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Rapid formation of secondary aerosol precursors from the autoxidation of C_5–C_8 *n*-aldehydes

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Abstract. Long chain aldehydes are common atmospheric constituents, and their gas-phase oxidation form low volatility condensable products leading to secondary organic aerosol. Although the oxidation of *n*-aldehydes initiated by OH radicals is dominated by aldehydic hydrogen abstraction, the non-aldehydic hydrogen abstractions tend to become competitive with the increase of aldehyde carbon chain length. Here, we experimentally investigated the oxidation of C_5–C_8 *n*-aldehydes in variable reaction times (1–13 s) in a flow tube reactor coupled to a nitrate ion time-of-flight chemical ionization mass spectrometer (NO_3^--ToF-CIMS). Octanal produced highly oxygenated organic molecules (HOMs – low volatility products) with up to 7 O atoms within 1.0 s while the same level of oxygenation was acquired by pentanal within 2.3 s. In long reaction time (11–13 s) experiments, we observed HOMs with progressively more O atoms and higher product yields with the increase of carbon atoms in the precursor aldehydes. Our experiments in the presence of high NO concentrations (2 ppb to 1 ppm) showed the formation of prominent highly oxygenated organonitrates along with the suppression of HOM accretion products. However, some enhancement in the monomeric HOMs even with 6 O atoms were seen under variable NO conditions. Results from hydrogen to deuterium (H/D) exchange experiments showed that the studied *n*-aldehydes undergo similar autoxidation mechanisms, but the reactivity and HOM formation potential increase with increasing carbon chain length.

1 Introduction

Secondary organic aerosol (SOA) refers to the aerosol material that is mainly formed by the atmospheric gas-phase oxidation of volatile organic compounds (VOCs) (Kroll and Seinfeld, 2008; Ziemann and Atkinson, 2012; Seinfeld and Pandis, 2016). Atmospheric oxidation of VOCs increases the oxygen to carbon ratios (O : C) in the oxidation products and can form highly oxygenated organic molecules (HOMs). These low volatility products are found to play a key role in the formation and growth of SOA (Ehn et al., 2014; Öström et al., 2017; Bianchi et al., 2019; Brean et al., 2019, 2020). However, heterogeneous and multiphase chemistry involving the reactions of organic compounds directly onto solid particles or inside liquid particles can also be an important contributor to SOA mass (Ervens et al., 2011, 2014; Kuang et al., 2020; Gu et al., 2023). The SOA is a dominant component of tropospheric fine particulate matter (Hallquist et al., 2009; Spracklen et al., 2011; Huang et al., 2014), influences oxidative capacity, local and global air quality, climate change, and human health (Jacobson et al., 2000; Hansen and Sato, 2001; Kanakidou et al., 2005; Zhang et al., 2014). Despite having significant attention due to its importance on SOA, its sources and formation processes are yet to be fully understood.

Aldehydes are common emissions in natural and polluted environments and have both biogenic and anthropogenic sources (Lipari et al., 1984; Carlier et al., 1986; Ciccioli et al., 1993; Schauer et al., 1999a, b, 2001) and are also formed by chemical transformation of other VOCs, especially ozonolysis of alkenes (Calogirou et al., 1999). At-

mospheric degradation of aldehydes is mainly governed by photolysis and the reaction with OH radicals in the daytime (Mellouki et al., 2003; Calvert et al., 2011; Mellouki et al., 2015). During nighttime, the reactions with NO_3 radicals are the dominant sink of aldehydes (Calvert et al., 2011). Although there are prior kinetic studies of reactions of *n*-aldehydes with OH radicals available in the literature (Albaladejo et al., 2002; Cassanelli et al., 2005; Iuga et al., 2010; Castañeda et al., 2012; Wang et al., 2015; Aguirre et al., 2025), they are limited to the initial steps of oxidation except recent studies (Barua et al., 2023; Yang et al., 2024) showing further oxidation steps leading to the formation of more functionalized products including HOMs.

It is well understood that the reaction of aldehydes with OH radicals is predominantly initiated by the abstraction of aldehydic hydrogen atom due to its weaker bond strength. However, with the increase of carbon chain length, the abstraction of other hydrogen atoms distant from the aldehydic moiety can also contribute to the overall oxidation process. The aldehydic H abstraction can lead to the cleavage of that carbon (C1) by CO loss from acyl (RC(O)) intermediate (Rissanen et al., 2014; Barua et al., 2023). Alternatively, it leads to the acyl peroxy (RC(O)OO) and, in high NO_x ($NO + NO_2$) condition, subsequently to acyloxy (RC(O)O) intermediate followed by CO_2 loss ultimately forming C_{n-1} aldehyde, C_{n-1} alkyl nitrate, and C_{n-1} alkoxy isomerization products (Vereecken and Peeters, 2009; Chacon-Madrid et al., 2010). Besides, it can also produce C_n peroxyacyl nitrates (PAN) via RC(O)OO + NO_2 reaction (Mellouki et al., 2003; Chacon-Madrid et al., 2010; Calvert et al., 2011; Mellouki et al., 2015), a reservoir species for long-range transport of NO_x in the free troposphere. Additionally, C_n peroxy acids can be formed by the reaction of RC(O)OO intermediate and HO_2 (Barua et al., 2023).

Chacon-Madrid et al. (2010) have conducted a comparative study of the SOA yields in OH-initiated oxidation of *n*-aldehydes and *n*-alkanes under high NO_x conditions. They reported near identical SOA yields from *n*-tridecanal and *n*-dodecane where the precursor alkane is having one less carbon than the precursor aldehyde. The finding was attributed to the formation of C_{12} alkoxy radical intermediate from both precursors undergoing similar subsequent reactions leading to SOA. This indicates that the dominant fate of aldehyde oxidation by OH leads to the fragmentation of its carbon backbone losing one carbon atom rather than producing functionalized products with the same number of carbon atoms as the parent molecule.

In low NO_x conditions, previous experimental studies have shown that the abstraction of the aldehydic hydrogen can also lead to molecular functionalization forming HOMs (Ehn et al., 2014; Rissanen et al., 2014; Tröstl et al., 2016; Wang et al., 2021) and potentially promote the SOA yields. Recently, Barua et al. (2023) have studied hexanal OH oxidation reaction in detail using high level quantum chemical computations as well as experimental mass spectrometry technique. They showed that both the aldehydic and non-aldehydic H abstraction pathways can contribute to the functionalization of hexanal resulting in rapid formation of HOMs via autoxidation. Autoxidation refers to chain radical processes, generally starting with an oxygen-centered radical that undergoes unimolecular isomerization reaction leading to a carbon-centered radical species whose dominant fate is to add additional molecular oxygen and thus increases product O : C ratios (Crounse et al., 2013; Jokinen et al., 2014; Rissanen et al., 2014; Berndt et al., 2015, 2016; Mentel et al., 2015; Rissanen et al., 2015). Along the aldehydic H abstraction reaction route, the fastest isomerization (1,6 H-shift rate coefficient, $k = 0.2\,s^{-1}$) of the RC(O)OO intermediate was shown to be the key for autoxidation reaction chain propagation and competitive with any bimolecular reaction mediated RC(O)O fragmentation (Barua et al., 2023). Thus, the RC(O)OO isomerization reaction (Seal et al., 2023) keeps the carbon backbone of the precursor aldehyde intact. A non-aldehydic H abstraction from C4 was also seen to be competitive and its corresponding RO_2 was shown to undergo a 1,6 H-shift reaction with the aldehydic H atom at a higher rate ($k = 0.9\,s^{-1}$) than the H-shifts in RC(O)OO radical. The fast aldehydic H-shift to the peroxy group is consistent with other carbonyl systems reported previously (Da Silva, 2011; Crounse et al., 2012; Møller et al., 2016, 2019). Moreover, the kinetic modelling simulation on OH-initiated oxidation of hexanal conducted by Barua et al. (2023) showed that a detectable concentration (1.3×10^4 molec. cm^{-3}) of O_7 HOMs were produced even in the presence of 1 ppb NO. Because the non-aldehydic H abstraction pathways are likely less prone to fragmentation and promote functionalization, the effect of carbon chain length of *n*-aldehydes on HOM yields is of great interest.

In this work, we experimentally studied the OH initiated autoxidation of C_5–C_8 *n*-aldehydes in variable reaction times under atmospheric pressure and room temperature using state-of-the-art mass spectrometry technique. In addition, the reactions were studied in the presence of variable concentrations of NO to examine the effect of NO on the oxidation process. The study shows how the length of carbon chain in linear aldehydes directly affects the reactivity and functionalization of the molecules during their oxidation initiated by OH radicals.

2 Methods

2.1 Experimental setup

The gas-phase oxidation reactions of *n*-aldehydes with OH radicals were conducted in a flow reactor setup in the laboratory as shown in Fig. 1. All the experiments were conducted at room temperature and 1 atm pressure of air. The precursor aldehydes were introduced into the reactor from their individual gas cylinders. The oxidant OH radicals were produced in situ by the reaction of tetramethylethylene

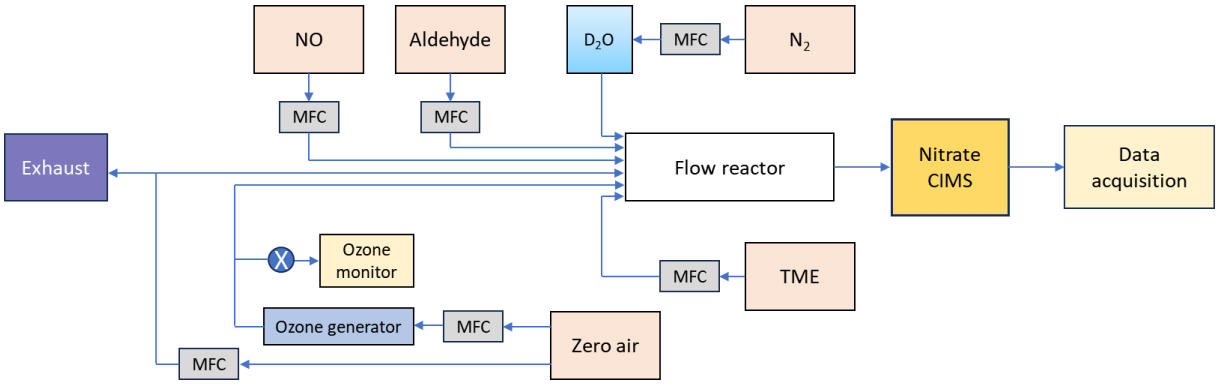

Figure 1. Schematic representation of a flow reactor setup showing a nitrate (NO_3^-) based chemical ionization mass spectrometer coupled to ambient pressure flow reactor. TME = tetramethylethylene (C_6H_{12}). The oxidant OH radical was produced in situ by TME + O_3 reaction. MFC = mass flow controller.

(TME, C_6H_{12}) with ozone. We used an ozone generator that photolyzes zero-air by a mercury lamp (UVP, Analytik Jena) to provide ozone to the reactor while TME was supplied from a gas cylinder. The zero-air was produced by feeding compressed clean air to a zero-air generator (AADCO-737-15) which was also used as bath gas in the reactor maintaining a total sample flow of 8–10 slpm. The initial concentrations of reactant precursors were determined by controlling the individual gas flows using calibrated mass flow controllers (Alicat Scientific). An ozone analyzer (2B Technologies model 205) was used to measure the ozone concentrations. Further details regarding reactant concentration measurements, as well as specifications for the chemicals and gas cylinders, are provided in Sects. S1 and S2 in the Supplement.

We studied the reactions over a range of reaction times (short: 1–3 s, and long: 11–13 s). The experimental conditions are presented in Table 1. Long reaction time experiments were conducted using a borosilicate flow reactor (length: 100 cm, and i.d.: 4.7 cm) while a quartz flow reactor (length: 100 cm, and i.d.: 2.2 cm) was used for the short reaction time experiments. We utilized the full volume of the reactor to achieve a long reaction time. However, the short reaction times were achieved by controlling the distance between the mass spectrometer orifice and the position where the precursor aldehyde meets the oxidant OH inside the reactor. This was done by using a movable injector that brings the aldehyde of interest at variable positions inside the reactor. The shortest possible reaction time of an individual aldehyde was chosen by the detection of any HOMs from its oxidation initiated by OH radicals. In short reaction time experiments, the highest concentration of VOC (6.4 ppmv) was used for pentanal while the concentrations of other aldehydes were up to 1 ppmv. The VOC concentrations of 0.2–2.5 ppmv were used in long reaction time experiments. The other reactants including 43–97 ppbv of TME, and 77–295 ppbv of ozone were maintained nearly constant with respect to individual VOC in all experiment types (see Table 1). Among all the studied *n*-aldehyde systems, the lowest level of aldehyde and O_3 concentrations were maintained for heptanal oxidation experiment. This is because higher concentrations led to irrelevant products likely originating from the ozonolysis of heptanal stabilizers (see Fig. S2). To observe the effect of NO_x on the OH induced oxidation of *n*-aldehydes, we conducted the experiments with variable NO concentrations (2–1000 ppbv). Additionally, one more set of experiments, hydrogen to deuterium (H/D) exchange, were conducted by the addition of D_2O flow to the *n*-aldehyde OH oxidation reaction. These experiments give an estimate of the number of functional groups with labile H atoms (e.g., OH, OOH, and C(O)OOH) in the oxidation products. A near complete H/D conversion was confirmed by monitoring the reagent ion signals of $HNO_3NO_3^-$ and $(HNO_3)_2NO_3^-$ fully converting to $DNO_3NO_3^-$ and $(DNO_3)_2NO_3^-$, respectively (see Fig. S14). The time series of reactive species, VOC, OH, RO_2, O_3, and NO under different reaction conditions are shown in Figs. S16 and S17.

The oxidation products were detected using a nitrate ion time-of-flight chemical ionization mass spectrometer (NO_3^--ToF-CIMS) as their NO_3^- adducts. A zero-air sheath flow of 20 slpm was provided to the chemical ionization inlet. The NO_3^- ions were produced from gas-phase HNO_3 flow under soft X-ray exposure while being carried by N_2 to the inlet. The mass spectrometric data processing, including averaging, baseline removal, mass axis calibration, and peak integration were done using the tofTools v6.03 package for MATLAB.

2.2 Kinetic simulation

We estimated the concentrations of reactive species in the flow reactor during different *n*-aldehyde oxidation experiments using a kinetic simulator Kinetiscope (version 1.1.1136.x64) (Hinsberg and Houle, 2022). These include the average concentrations of OH radicals and initial RO_2 radicals (i.e., $C_nH_{2n-1}O_3$, the first peroxy radicals formed in

Table 1. The experimental conditions for OH initiated oxidation of studied C_5–C_8 *n*-aldehydes.

Experiment type VOC	[VOC][a] ppmv	[TME][a] ppbv	[O_3][a] ppbv	[OH][b] pptv	[NO][a] ppbv	D_2O y/n[c]	Δt[d] s
Short residence time							
Pentanal ($C_5H_{10}O$)	6.4	48.2	295	4.4	n/a	n	2.3
Hexanal ($C_6H_{12}O$)	1.0	43.2	225	3.4	n/a	n	1.1, 2.9
Octanal ($C_8H_{16}O$)	0.72	48.2	208	3.1	n/a	n	1.0, 2.1
Long residence time							
Pentanal ($C_5H_{10}O$)	2.5	48.2	295	4.4	n/a	n	12.8
Hexanal ($C_6H_{12}O$)	1.0	43.2	225	3.4	n/a	n	11.5
Heptanal ($C_7H_{14}O$)	0.15	96.5	77	1.2	n/a	n	12.8
Octanal ($C_8H_{16}O$)	0.72	48.2	208	3.1	n/a	n	12.8
Experiments with NO							
Pentanal ($C_5H_{10}O$)	1.3	48.2	208	3.1	2–1000	n	12.8
Hexanal ($C_6H_{12}O$)	1.0	43.2	225	3.4	2–200	n	11.5
Octanal ($C_8H_{16}O$)	0.72	48.2	208	3.1	2–1000	n	12.8
Experiments with D_2O							
Pentanal ($C_5H_{10}O$)	2.5	48.2	295	4.4	n/a	y	12.8
Hexanal ($C_6H_{12}O$)	1.0	43.2	225	3.4	n/a	y	11.5
Heptanal ($C_7H_{14}O$)	0.10	96.5	77	1.2	n/a	y	12.8
Octanal ($C_8H_{16}O$)	0.72	48.2	208	3.1	n/a	y	12.8

[a] Initial reactant concentrations. [b] The initial OH concentrations were calculated using bimolecular rate coefficients $k_{O_3-TME} = 1.5 \times 10^{-15}$ cm^3 molec.$^{-1}$ s^{-1}, $k_{OH-TME} = 1.0 \times 10^{-10}$ cm^3 molec.$^{-1}$ s^{-1} (Manion et al., 2015), and the expression $[OH] = (k_{O_3-TME} \times [O_3])/k_{OH-TME}$. n/a = not applicable. [c] D_2O added = y, not added = n. [d] Reaction time (Δt).

the oxidation process) produced in the experiments. In the simulations, a single-reactor model with constant volume, pressure, and temperature was employed. The temperature was set to 298.15 K. The simulation setting parameters such as total number of particles (1×10^9) and random number seed (12947) were kept constant for consistency, while the maximum simulation time was matched to that of individual experiments. Details of all the simulations are provided in Sect. S13.

3 Results and discussion

3.1 Detection of HOM in short to long reaction time experiments

In this section, we discuss how early HOMs formed and how they evolved with the progress of reaction time in different *n*-aldehyde OH oxidation experiments. The hexanal OH oxidation spectra included in Figs. 2 and 3 are reproduced from our previous study (Barua et al., 2023) for comparison with the other aldehydes. Figure 2 shows the results from short reaction time ($\Delta t = 1$–3 s) experiments. We observed the formation of O_6 and O_7 HOMs within 1.0, 1.1, and 2.3 s reaction times in the oxidation experiments of octanal, hexanal, and pentanal, respectively. It is essential to mention that with the increase of number of carbon atoms in the studied aldehydes, the required precursor concentrations for first HOM observation decreased (from 6.4 ppm pentanal, 1 ppm hexanal to 0.72 ppm octanal; corresponding reacted concentrations 1.15, 0.34, and 0.30 ppb, respectively). This shows a clear effect of carbon chain length on the reactivity of linear aldehydes towards HOM formation. Because the oxidation process of *n*-aldehydes ($C_nH_{2n}O$) with OH is initiated by the abstraction of a H atom (aldehydic or non-aldehydic), the first formed acyl (or alkyl) peroxy radical $C_nH_{2n-1}O_3$ contains an odd number of oxygen atoms. If autoxidation outcompetes any other bimolecular reactions (e.g., $RO_2 + RO_2$, $RO_2 + HO_2$, etc.), the product spectrum will be mostly dominated by odd number of oxygen containing products. In all studied *n*-aldehyde systems, the intensity of O_6 HOM is higher than that of O_7 HOM (Figs. 2 and 3). The formation of $C_nH_{2n-1}O_6$ peroxy radical indicates that the process certainly involves a bimolecular reaction step. In hexanal OH oxidation, Barua et al. (2023) computationally showed that the formation of $C_6H_{11}O_5$ peroxy radical via autoxidation is very fast while the subsequent isomerization reaction leading to $C_6H_{11}O_7$ is slower. The same is likely to hold true for other aldehydes and it is expected that the $C_nH_{2n-1}O_5$ peroxy radical undergoes a bimolecular reaction converting it to $C_nH_{2n-1}O_4$ alkoxy radical, followed by a H-shift, and subsequent O_2 addition reactions producing the dominant $C_nH_{2n-1}O_6$ HOM (see Figs. S7–S11). As the reaction time increased, we observed the formation of monomeric

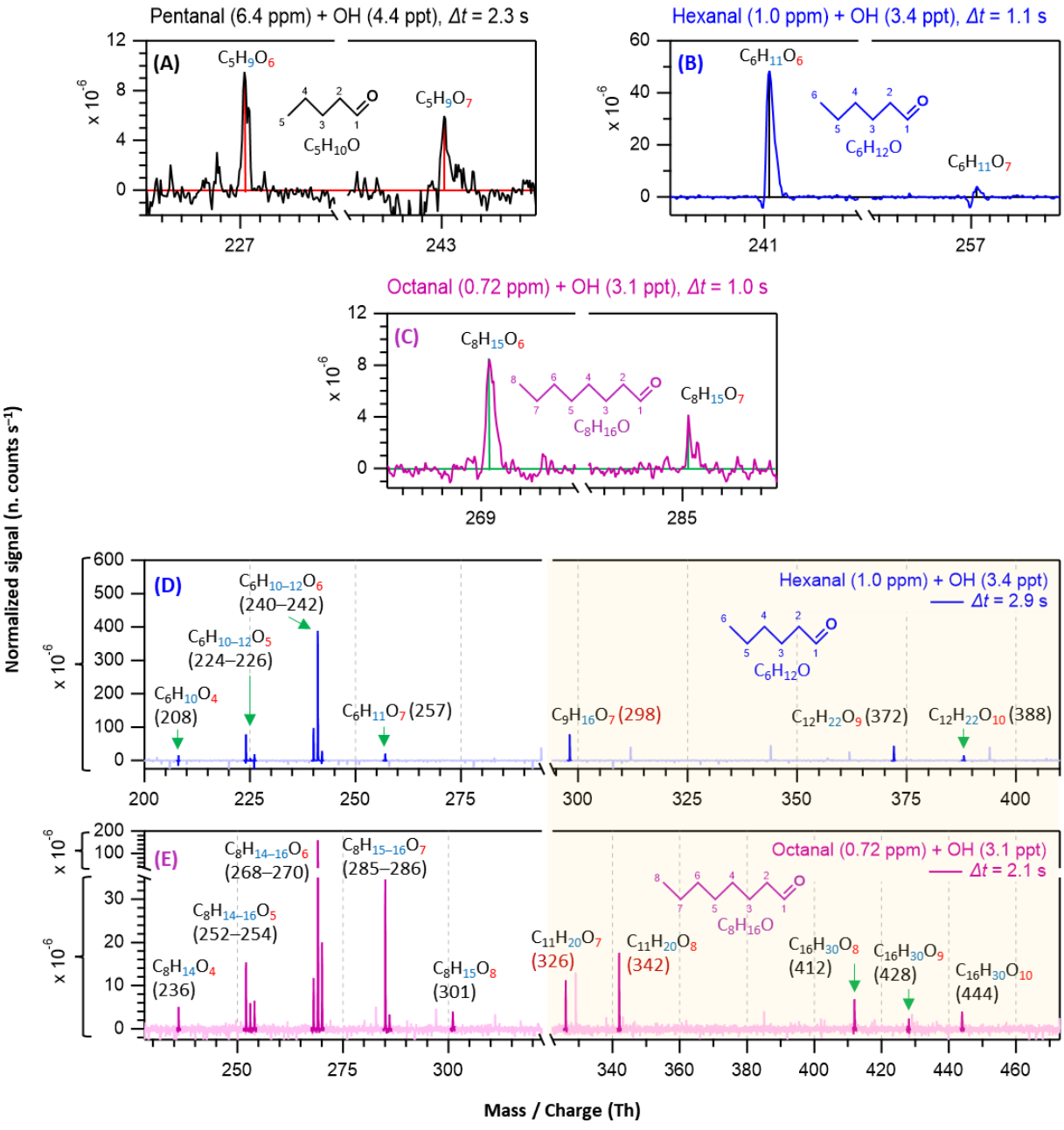

Figure 2. Nitrate chemical ionization mass spectra of OH initiated oxidation of *n*-aldehydes showing the formation of HOMs in different reaction times (Δt): 2.3 s – pentanal in black **(a)**, 1.1 and 2.9 s – hexanal in blue **(b, d)**, and 1.0 and 2.1 s – octanal in purple **(c, e)**. The product peaks are labelled with the exclusion of NO_3^- ion attachment in their compositions. The backgrounds of TME ozonolysis (TME + O_3) and aldehyde have been subtracted from all spectra, resulting in several negative peaks in panels **(a–e)**. The accretion product region is highlighted in light gold background. The accretion products labeled with nominal mass/charge in dark red ($C_{n+3}H_{2n+4}O_{7-8}$) are related to the TME-derived peroxy radical $C_3H_5O_3$.

HOM up to O_7 and accretion products up to O_{10} composition within 2.9 s in hexanal oxidation (see Fig. 2d) with the consumption of $\sim 8.88 \times 10^{-4}$ of its initial concentration. In the case of octanal, monomeric O_8 HOM and accretion products up to O_{10} formed within 2.1 s reaction time (see Fig. 2e) with a comparable reacted fraction ($\sim 8.71 \times 10^{-4}$) of its initial concentration.

In long reaction time (11–13 s) experiments, we observed higher intensities of product signals (see Fig. 3) in comparison to their intensities in the short reaction time experiments – as expected. For pentanal, Fig. 3a shows that

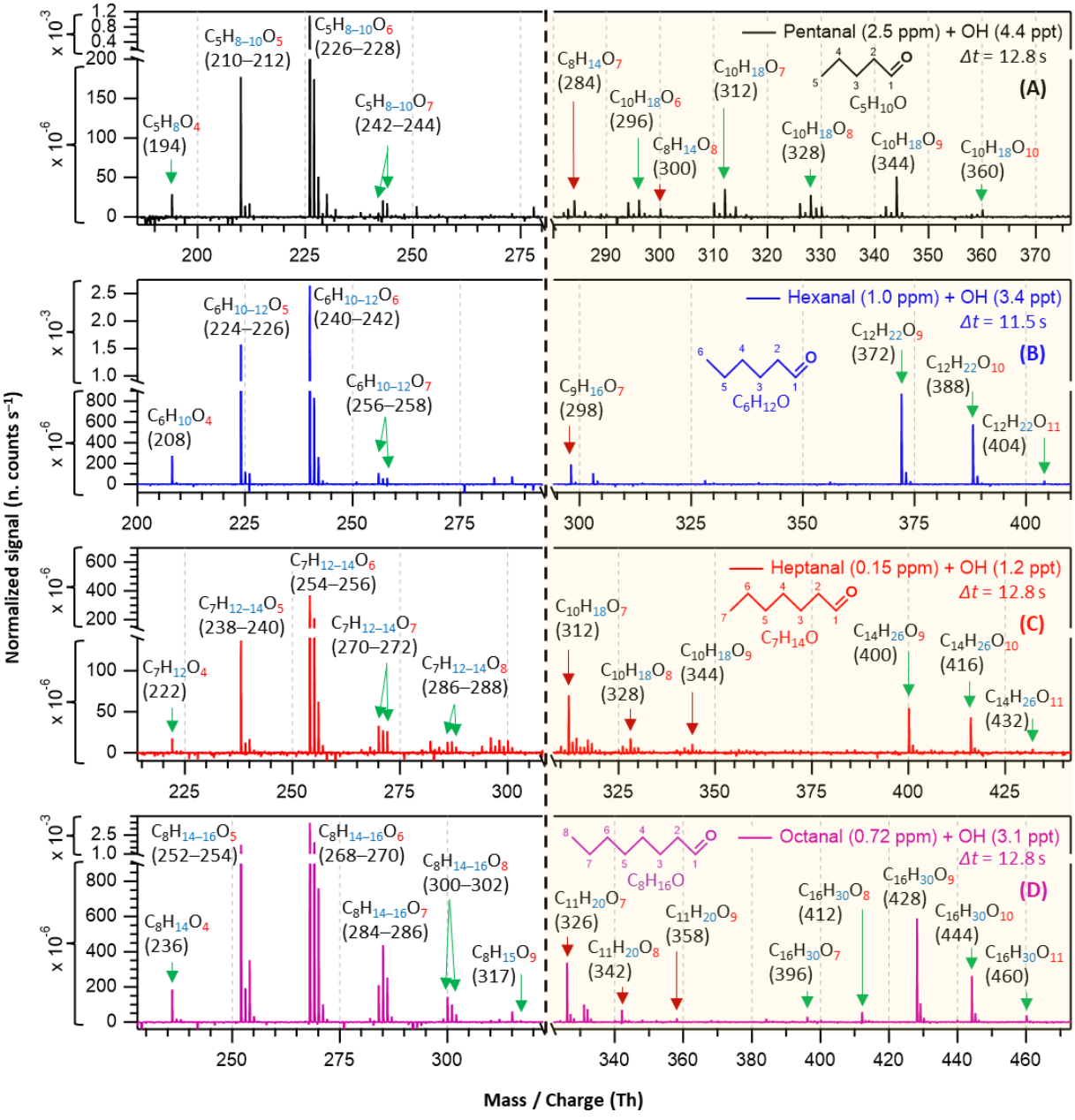

Figure 3. Nitrate chemical ionization mass spectra of OH initiated oxidation of *n*-aldehydes in 11–13 s residence times: pentanal in black **(a)**, hexanal in blue **(b)**, heptanal in red **(c)**, and octanal in purple **(d)**. The product peaks are labelled with the exclusion of NO_3^- ion attachment in their compositions. The backgrounds of TME ozonolysis (TME + O_3) and aldehyde have been subtracted from all spectra, resulting in several negative peaks in panels **(a–d)**. The accretion product region is highlighted in light gold background. The accretion products marked by dark red arrows ($C_{n+3}H_{2n+4}O_{7-9}$) are related to the TME-derived peroxy radical $C_3H_5O_3$.

HOM accretion products up to O_{10} formed within 12.8 s reaction time which were not seen in the short reaction time experiment (2.3 s). A lower precursor concentration, 2.5 ppm of pentanal in long reaction time experiment compared to earlier 6.4 ppm in short reaction time, was sufficient to produce the observed HOMs in this case. A close observation of C_6–C_8 *n*-aldehyde oxidation spectra (Fig. 3b–d) reveals that HOM accretion products up to O_{11} formed within 11–13 s reaction time under the experimental conditions. In all *n*-aldehyde experiments, we also observed the accretion products resulting from different combinations of aldehyde-derived peroxy radicals $C_nH_{2n-1}O_{6-8}$ and TME-derived peroxy radical $C_3H_5O_3$ (see Fig. S13 for details) which are marked with dark red arrows. Figure 3d implies that the highest oxygenation ($C_8H_{15}O_9$) in the monomeric HOM products is associated with octanal, whereas Fig. 3c

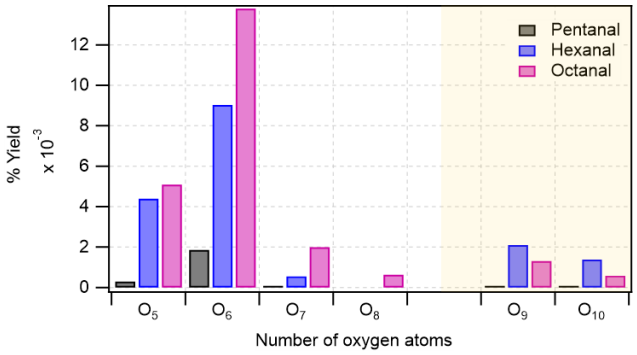

Figure 4. Distribution of major oxidation products (O_5–O_8 in monomeric regime with white background, and O_9–O_{10} in accretion product regime with light gold background) in pentanal, hexanal, and octanal oxidation initiated by OH radical. In *y*-axis, the numbers are the cumulative sum of yields of products with the same oxygen number. Reaction time, $\Delta t = 11$–13 s.

shows that the most oxygenated products produced from heptanal are $C_7H_{12-14}O_8$. In the case of pentanal and hexanal, monomeric HOMs are limited to seven oxygen atoms (see Fig. 3a and b). All in all, we notice a near identical distribution of oxidation products in the experiments with all C_5–C_8 *n*-aldehydes. However, the tendency of oxidation gets faster and advances to higher oxygenated products when the carbon chain length increases.

The bar plot (see Fig. 4) compares the yields of major oxidation products from different *n*-aldehydes. It clearly shows that the yields of higher oxygenated products increase as we move from pentanal to octanal. In the accretion product regime, the yields of O_9–O_{10} products in octanal are lower than that of hexanal which is also reflected in their dimer to monomer ratios with octanal being 8.8×10^{-2} and hexanal being 2.5×10^{-1}. The lower ratio for octanal compared to hexanal is observed despite both precursors producing comparable quantities of initial RO_2 radicals (9.12×10^9 and 9.37×10^9 molec. cm^{-3} from octanal and hexanal, respectively; see Table S1 in the Supplement). This can lie in the variation of $RO_2 + RO_2$ reaction rate coefficients forming the accretion products ($RO_2 + RO_2 \rightarrow ROOR + O_2$) which is highly dependent on specific RO_2 structures (Shallcross et al., 2005; Berndt et al., 2018).

3.2 Experiments in the presence of NO

It has been widely acknowledged that the formation of HOM is suppressed in high NO_x conditions (Wildt et al., 2014; Praske et al., 2018; McFiggans et al., 2019; Pullinen et al., 2020), thus reducing the SOA yields. In this process, the reduction in SOA yield is largely attributed to the suppression of highly condensable HOM accretion products ($RO_2 + RO_2 \rightarrow ROOR$) (Kirkby et al., 2016; Pullinen et al., 2020). However, other studies have shown that NO can also enhance HOM formation by producing reactive RO radicals ($RO_2 + NO \rightarrow RO + NO_2$) that can propagate autoxidation (Rissanen, 2018; Yan et al., 2020; Wang et al., 2021; Shen et al., 2022; Nie et al., 2023; Barua et al., 2025; Kang et al., 2025). In our different *n*-aldehyde oxidation experiments in the presence of variable concentrations of NO, we observed the general tendency of dropping of the accretion product signals ($C_{2n}H_{4n-2}O_z$) and their corresponding yields as expected (see Figs. 5 and 6). Figure 5 represents the mass spectra recorded in the experiments without (in blue) and with the presence of 100 ppb initial NO (in red). At this condition, most of the monomeric products are quenched at different extents alongside the formation of organonitrates while the accretion products are quenched nearly completely. Interestingly, we observed some enhancement in the intensities of O_4 products in the case of hexanal and octanal (Fig. 5b and c) under 100 ppb NO. A closer look at Fig. 6a reveals that the yields of closed-shell products $C_5H_8O_{4-6}$ somewhat increased under around 1 ppb average NO condition which then started to decrease under higher NO of around 30 ppb and above in pentanal oxidation. In both pentanal and hexanal oxidations, the dominant O_6 peroxy radicals ($C_5H_9O_6$ and $C_6H_{11}O_6$, respectively, blue markers) gained higher yields under around 30 ppb of NO (Fig. 6a and b) compared to without NO condition.

On the other hand, in octanal oxidation, the yield of dominant O_6 peroxy radical ($C_8H_{15}O_6$) increased even under 70 ppb of NO (see Fig. 6c). These observations indicate that the suppressing effect of NO on the yields of HOMs in *n*-aldehyde oxidation is perturbed with the increase of carbon chain length of the precursor aldehyde. Also, in all the studied *n*-aldehydes, although highly oxygenated products are suppressed under higher NO concentrations, we see some enhancement in the early oxygenated closed shell products (O_4–O_5, and even O_6 product in pentanal) under relatively lower NO concentrations. It should be noted that the formation of organonitrates with chemical composition $C_nH_{2n-1}O_zNO$ strongly supports our assignment of reactive peroxy radical intermediates $C_nH_{2n-1}O_z$. Additional insights into the abundance of HOMs under varying NO concentrations are discussed in Sect. S12.

3.3 D_2O experiments

With the addition of D_2O in the OH initiated oxidation experiments of *n*-aldehydes, we observed a shift in individual product signals in the mass spectra equivalent to the number of exchangeable H atoms in the product structures (see Fig. 7). This provides additional insight into the product identities in terms of the total number of OH, OOH, and (or) C(O)OOH groups present in their molecular structures. Figure 7 shows that the oxidation products with same number of oxygen atoms in all the studied *n*-aldehydes undergo identical mass shifts (H/D exchange) in the presence of D_2O. This observation indicates that the autoxidation mechanism derived by Barua et al. (2023) for hexanal OH oxidation (see

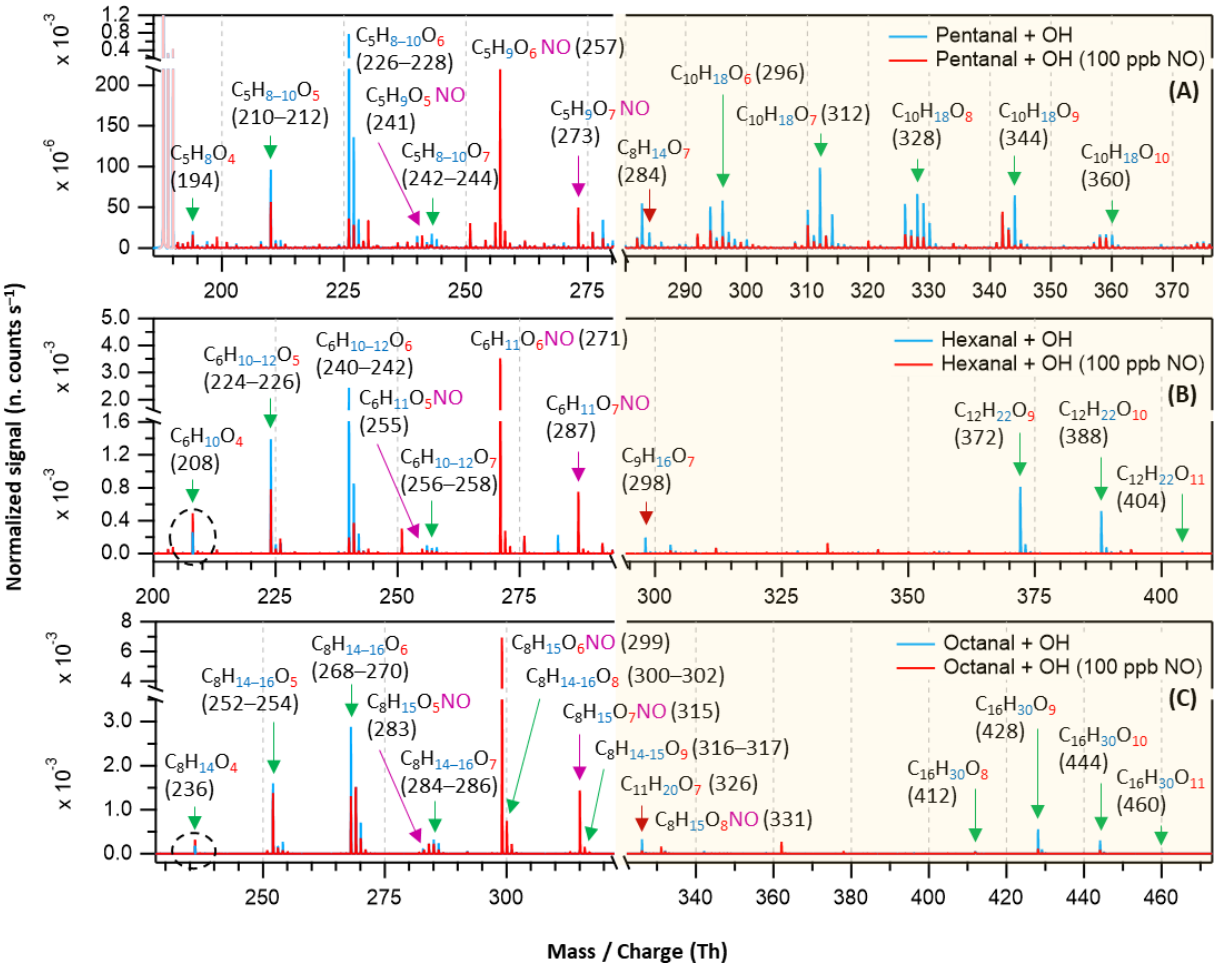

Figure 5. Overlaid nitrate chemical ionization mass spectra of OH initiated oxidation of *n*-aldehydes without (in blue) and with the presence of 100 ppb of initial NO (in red): pentanal **(a)**, hexanal **(b)**, and octanal **(c)**. The product peaks are labelled with the exclusion of NO_3^- ion attachment in their compositions. The accretion product region is highlighted in light gold background. The accretion products marked by dark red arrows ($C_{n+3}H_{2n+4}O_{7-9}$) are related to the TME-derived peroxy radical $C_3H_5O_3$. The organonitrates (in the presence of NO) are labelled with the extension NO and marked with purple arrows. Reaction time, $\Delta t = 11$–13 s.

Fig. S5) is directly applicable to the other linear aldehydes studied here and thus produces similar product structures. The original mechanism is extended to HOMs up to nine oxygen atoms and presented in Fig. S6. The likely formation process of HOM accretion products ($C_{2n}H_{4n-2}O_{9-11}$) is shown in Fig. S12. The proposed structures of O_5 closed-shell products, O_6–O_8 monomeric HOMs as well as O_{9-11} HOM accretion products (see Figs. S6–S12) agree with the H/D exchange experiments in terms of their 2–4 units of mass shifts in the respective *n*-aldehyde mass spectra (see Fig. 7).

3.4 Atmospheric implications

Ambient concentration of individual longer chain aldehyde ($\geq C_5$) can vary from sub-ppb to several ppb depending on time and location (Williams et al., 1996; Duan et al., 2008; Li et al., 2018; Ma et al., 2019). A total concentration of C_6–C_{10} *n*-aldehydes in Monti Cimini Forest in Italy was measured to be 8.8 ppb (Ciccioli et al., 1993). In indoor air, the concentration can be significantly higher, even around 50 ppb (Birmili et al., 2022). Atmospheric lifetime of *n*-aldehydes due to their reactivity with OH radicals is generally less than 10 h (Albaladejo et al., 2002; Aguirre et al., 2025). Because of their significant photochemical ozone formation potential (Jenkin et al., 2017; Aguirre et al., 2025), they are good candidates for generating photochemical smog in NO_x rich polluted urban atmosphere. On the other hand, previous studies have shown that atmospheric oxidation products of longer chain *n*-aldehydes are direct contributors to the formation of SOA (Chacon-Madrid et al., 2010; Fan et al., 2024). Here, we demonstrated that the studied C_5–C_8 *n*-aldehydes can rapidly form HOM via autoxidation initiated by OH radicals, and the length of carbon chain controls the efficiency

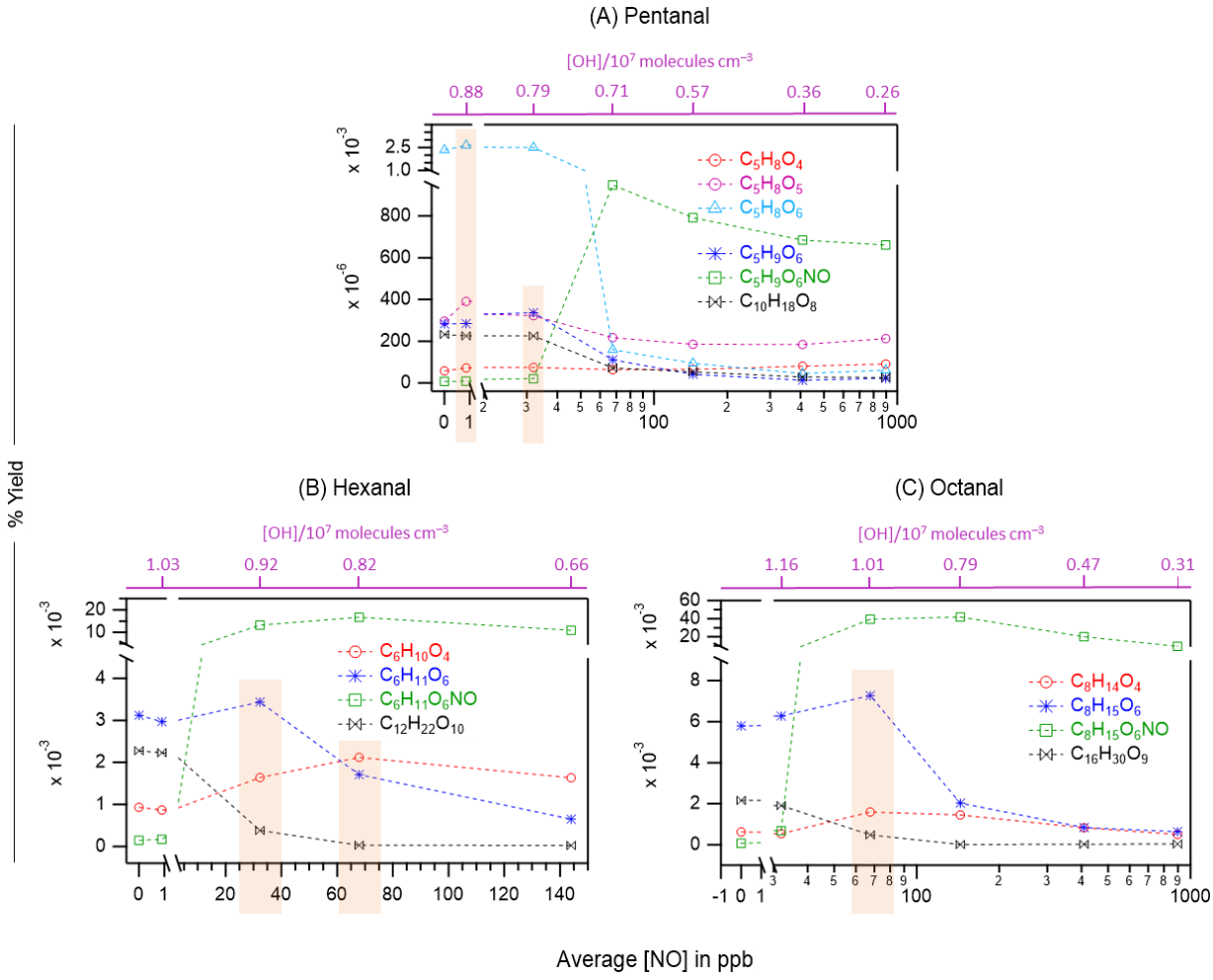

Figure 6. The yields of different oxidation products including monomeric HOMs, organonitrates (green markers), and HOM accretion products (black markers) as a function of average NO concentrations in OH initiated oxidation of *n*-aldehydes: pentanal **(a)**, hexanal **(b)**, and octanal **(c)**. The corresponding average OH radical concentrations in these experiments are shown with purple scales. Note the logarithmic scale (*x*-axis) in panels **(a)** and **(c)**. The orange rectangles highlight the enhanced yields of several non-nitrogen containing products: O_4–O_6 closed-shell products from pentanal under around 1 ppb NO **(a)**, O_4 closed-shell products from hexanal and octanal under around 70 ppb NO **(b, c)**, O_6 peroxy radical from pentanal and hexanal under around 30 ppb NO **(a, b)** and the same from octanal under 70 ppb NO **(c)**. Reaction time, $\Delta t = 11$–13 s.

of the process. Therefore, with the increase of carbon chain length in *n*-aldehydes, the fast formation of HOMs is expected to take part in the early stages of gas-to-particle formation and growth contributing to atmospheric SOA. Our experiments in the presence of NO showed a general decreasing trend of HOM accretion products with increasing NO but also formed the corresponding highly oxygenated organic nitrates (HOM-ONs). In our study with the *n*-aldehydes under 1–70 ppb NO conditions, some yield enhancement with oxidation products up to 6 O atoms was also seen. Pullinen et al. (2020) showed that both HOMs and HOM-ONs originated from the same peroxy radicals with more than six O atoms condensed on particles by about 50 % and those with more than eight O atoms condensed by about 100 % to form SOA in monoterpene photooxidation. Moreover, other reports show that HOM-ONs originated from different VOCs can contribute to low volatility products (Barua et al., 2025) and thereby particle growth and aerosol mass loading (Fry et al., 2014; Lee et al., 2016; Huang et al., 2019).

4 Conclusions

This study represents the significance of longer chain linear aldehydes ($\geq C_5$), key components of atmospherically abundant oxygenated volatile organic compounds (OVOCs), in rapid formation of HOMs upon atmospheric oxidation and their potential contribution to atmospheric SOA. Among the studied C_5–C_8 *n*-aldehydes, the fastest HOM formation is associated with octanal forming O_6 and O_7 HOMs within 1.0 s reaction time with low precursor loading. Pentanal and

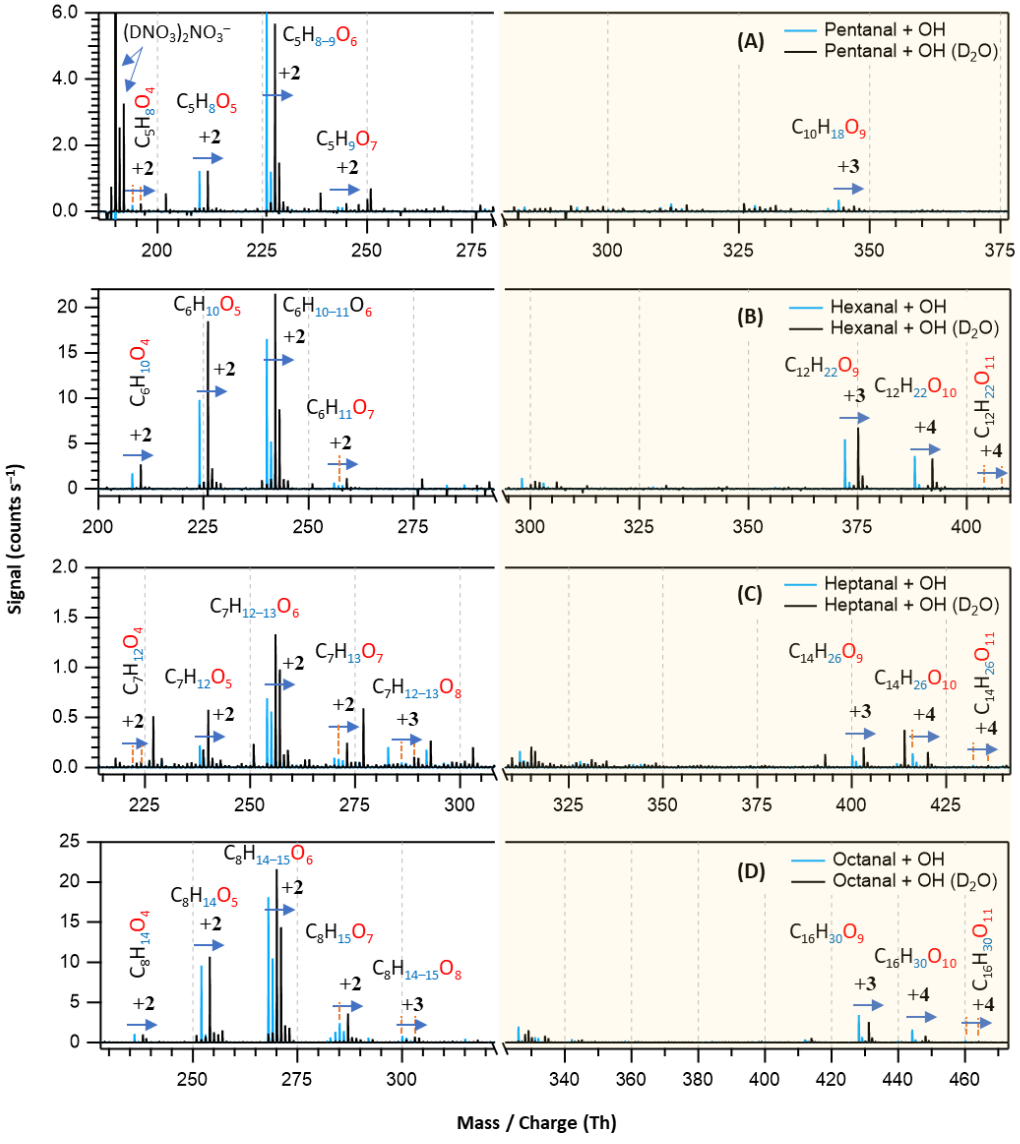

Figure 7. Overlaid nitrate chemical ionization mass spectra of OH initiated oxidation of *n*-aldehydes without (in blue) and with the presence of D_2O (in black): pentanal **(a)**, hexanal **(b)**, heptanal **(c)**, and octanal **(d)**. In panel **(a)**, the label $(DNO_3)_2NO_3^-$ is assigned to the deuterated nitric acid trimer reagent signal. The product peaks are labelled with the exclusion of NO_3^- ion attachment in their compositions and the numbers on the blue arrows indicate the individual counts of mass shift during H/D exchange. The backgrounds of TME ozonolysis (TME + O_3) and aldehyde have been subtracted from the spectra, resulting in several negative peaks in panels **(a)** and **(b)**. The accretion product region is highlighted in light gold background.

hexanal formed HOMs with the same number of oxygen atoms as early as 2.3 and 1.1 s, respectively but with higher precursor loadings (i.e., 6.4 ppm pentanal and 1.0 ppm hexanal compared to 0.72 ppm octanal). The highest oxygenated monomeric HOM with 9 oxygen atoms was formed from octanal whereas the numbers are up to 8 oxygen atoms for heptanal and 7 oxygen atoms for both pentanal and hexanal oxidation initiated by OH radicals within 13 s reaction time. Although the highest precursor concentration (6.4 ppm) was required for the first observation of detectable HOM in short reaction time experiment with pentanal, a lower concentration (2.5 ppm) was used to obtain its observed mass spectrum in the long reaction experiment. The HOM accretion products with up to 11 oxygen atoms were observed in C_6–C_8 *n*-aldehyde oxidation experiments while they were limited to maximum 10 oxygen atoms in pentanal case. We also observed the trend of increased oxidation product yields with the increase of carbon chain length. In all studied systems, the dominant product signals are O_6 HOMs with the O_5 HOMs being the second dominant ones. Previous mecha-

nistic understanding (Barua et al., 2023) as well as current experimental observations reveal that autoxidation process forming O_5 RO_2 in *n*-aldehydes is very fast while the subsequent unimolecular rearrangements of the RO_2 intermediates are in competition with bimolecular reactions including other RO_2, HO_2, and NO_x. The experiments in the presence of high NO concentrations (30 ppb and above) produced the highest yields of HOM-ONs, compared to neighboring non-nitrogen HOMs, with the expense of HOM accretion products. However, some enhancements with the yields of low oxygenated closed-shell products and O_6 peroxy radicals were also seen under 1–70 ppb NO conditions. The results of hydrogen to deuterium (H/D) exchange experiments with identical mass shifts in the oxidation products of all studied *n*-aldehydes imply that the autoxidation mechanism established for hexanal (Barua et al., 2023) is valid for other *n*-aldehydes. Therefore, accounting for linear aldehydes and their atmospheric oxidation with increasing importance to longer carbon chain length as a direct source of condensable materials even under moderately polluted urban areas is essential.

Data availability. Details about the experimental setup and mass spectrometry, chemicals and gas cylinders, mechanistic details of the oxidation steps, additional insights into abundance of HOMs, and details of kinetic simulations are provided in the Supplement. An Excel file (.xlsx) with results from mass spectrometry, and two example Kinetiscope (.rxn) files containing kinetic simulation input parameters – one with NO, another without – are available online (https://doi.org/10.5281/zenodo.18894230; Barua, 2026).

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Author contributions. Conceptualization: MR, SB, AK; data curation: SB, AK; formal analysis: SB; investigation: SB, AK, PS, SI, MR; methodology: SB, AK; writing (original draft preparation): SB; writing (review and editing): SB, AK, PS, SI, MR; funding acquisition: SI, MR.

Competing interests. The contact author has declared that none of the authors has any competing interests.

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