

# Deposited PtGe Clusters as Active and Durable Catalysts for CO Oxidation\*\*

Andoni Ugartemendia,<sup>[a]</sup> Jose M. Mercero,<sup>\*,[a]</sup> Abel de Cózar,<sup>[b, c]</sup> Marko M. Melander,<sup>[d]</sup> Jaakko Akola,<sup>\*,[e, f]</sup> and Elisa Jimenez-Izal<sup>\*,[a, c]</sup>

Control of CO emissions raises serious environmental concerns in the current chemical industry, as well as in nascent technologies based on hydrogen such as electrolyzers and fuel cells. As for now, Pt remains one of the state-of-the-art catalysts for the CO oxidation reaction, but unfortunately, it suffers from CO self-poisoning. Recently, Pt–Ge alloys were proposed to be an excellent alternative to reduce CO poisoning. This work investigates the impact of Ge content on the CO oxidation kinetics of Pt<sub>4</sub>Ge<sub>n</sub> subnanoclusters supported on MgO. A Ge concentration dependence of the reaction kinetics is found due to a strong synergy between Pt and Ge. Pt–Ge nanoalloys act as a bifunctional catalyst by displaying dual adsorption sites; i.e.,

CO is adsorbed on Pt whereas oxygen binds to Ge, forming an alternative oxygen source GeO<sub>x</sub>. Besides, Ge alloying modifies the electronic structure of Pt (ligand effects) and reduces the affinity to CO. In this way, the competition between CO and O<sub>2</sub> adsorption and the overbinding of CO is alleviated, achieving a CO poisoning-free kinetic regime. Our calculations suggest that Pt<sub>4</sub>Ge<sub>3</sub> is the optimal catalyst, evidencing that alloying composition is a parameter of extreme importance in nanocatalyst design. The work relies on global optimization search techniques to determine the accessibility of multiple structures at different conditions, mechanistic studies and microkinetic modeling.

## Introduction

Carbon monoxide (CO) is an odorless gas which is extremely hazardous for health and the environment. It is usually produced as a result of the incomplete combustion of hydrocarbon compounds. Thus the catalytic oxidation of CO has become a crucial step for the control of CO emissions in a wide range of processes, such as vehicle exhaust treatment and purification of hydrogen for fuel cell applications. Moreover, due to its simplicity, CO oxidation is used as a probe reaction to model more complex reactions such as the Fischer-Tropsch synthesis.<sup>[1–3]</sup>

Platinum group metals (PGMs) including Pt, Pd, Rh, etc. are widely known for their excellent catalytic activity for many reactions of critical importance such as CO oxidation.<sup>[4]</sup> Catalyst nanoparticles (NPs) are usually stabilized by dispersion over metal-oxide supports in industrial applications.<sup>[5]</sup> CO oxidation usually displays bistable kinetics over Pt.<sup>[6]</sup> At low temperature and/or low O<sub>2</sub> pressure, it suffers from CO self-poisoning (chemical deactivation), slowing down the reaction rate. The catalyst surface is then saturated with a CO layer that prevents the coadsorption of O<sub>2</sub>. Therefore, there are not enough free active sites in which the O<sub>2</sub> molecules can impinge and react with the coadsorbed CO, and the catalyst is rendered totally

[a] A. Ugartemendia, Prof. J. M. Mercero, Prof. E. Jimenez-Izal  
 Polimero eta Material Aurreratuak: Fisika, Kimika eta Teknologia Saila  
 Kimika Fakultatea  
 Euskal Herriko Unibertsitatea (UPV/EHU) and Donostia International Physics Center (DIPC)  
 20018 Donostia, Euskadi (Spain)  
 E-mail: jm.mercero@ehu.es  
 elisa.jimenez@ehu.es

[b] Prof. A. de Cózar  
 Kimika Organiko I Saila  
 Kimika Fakultatea  
 Euskal Herriko Unibertsitatea (UPV/EHU) and Donostia International Physics Center (DIPC)  
 20018 Donostia, Euskadi (Spain)

[c] Prof. A. de Cózar, Prof. E. Jimenez-Izal  
 IKERBASQUE  
 Basque Foundation for Science  
 Bilbao, 48009 Euskadi (Spain)

[d] Dr. M. M. Melander  
 Department of Chemistry  
 Nanoscience Center  
 University of Jyväskylä  
 FI-40014 Jyväskylä (Finland)

[e] Prof. J. Akola  
 Department of Physics  
 Norwegian University of Science and Technology (NTNU)  
 NO-7491 Trondheim (Norway)  
 E-mail: jaakko.akola@ntnu.no

[f] Prof. J. Akola  
 Computational Physics Laboratory  
 Faculty of Engineering and Natural Sciences  
 Tampere University  
 FI-33014 Tampere (Finland)

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inactive after solely a few number of catalytic cycles, costing billions of dollars annually in the industry.<sup>[7,8]</sup> At high temperature and/or high O<sub>2</sub> pressure on the other hand, a first order kinetic phase transition (*i.e.*, alteration of the adsorbates adsorption-desorption equilibrium) occurs leading to a more reactive state. Now, the adsorbed CO reacts immediately increasing the CO<sub>2</sub> formation rate and dissociated O becomes the predominant species. Since high temperature favors catalyst deactivation due to sintering and coking, CO poisoning imposes severe constraints over the maneuverability of the optimal operating reaction conditions. This scenario together with the high cost and limited availability of Pt hampers its use in large-scale production.

Subnanometer clusters have attracted a widespread attention in the last decades since they show unique catalytic properties compared with larger NPs due to finite-size effects.<sup>[9–12]</sup> In this regime the catalytic properties become size dependent to the extreme that the addition or subtraction of a single atom can have a drastic impact on them; this is the regime where each atom counts.<sup>[13,14]</sup> More importantly, they possess a huge surface to volume ratio which allows a maximal utilization of atoms, *i.e.*, the so-called atomic efficiency. It is important to note that, treating clusters as static structures has been shown to be a simplified picture.<sup>[15–17]</sup> In reality, clusters display high fluxionality at usual operating temperatures and form an ensemble of statistically relevant isomers, each of which exhibits different catalytic properties and constantly interconverting in response to external stimuli. All in all, clusters present a very strong structure-activity relationship by which the catalytic properties can be tuned by physical and chemical modifications (*e.g.*, size, composition, support, ligand capping, etc.); in short, clusters offer a myriad of possibilities for nano-catalyst design.

Alloying the catalyst with an ancillary element is a promising strategy to mitigate CO poisoning.<sup>[18]</sup> One of the advantages of this strategy is that it allows a reduction in the Pt loading, making them more attractive for commercialization. It is desirable that the bimetallic catalyst should maintain at least the same catalytic activity as Pt, as well as improve the selectivity towards different poisons. Several Pt-based bimetallic systems have been found in the last years for CO oxidation. For instance, Pt–Ni/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub>,<sup>[19]</sup> Pt–Fe/GA,<sup>[20]</sup> Pt–Fe/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub>,<sup>[21]</sup> Pt–Sn/AC,<sup>[22,23]</sup> Pt–Ru/SiO<sub>2</sub><sup>[24]</sup> and Pt–Re/SiO<sub>2</sub><sup>[25]</sup> outperform their monometallic counterparts with higher CO conversion rates. The enhancement of the bimetallic catalysts is usually ascribed to ligand (electronic) and/or bifunctional effects. Ligand effects occur when the alloying element modifies the Pt electronic structure (*e.g.*, electron donation and back-donation capacity of Pt) without directly participating in the reaction, while in bifunctional effects the promoter becomes an active component either by forming new bimetallic active ensemble sites or offering new single active sites with different adsorption properties (dual adsorption). This dual adsorption can be achieved by incorporating oxophilic promoters into Pt so that the oxygenated species now adsorb preferentially on them, leaving the Pt sites free for CO adsorption and thus, facilitating CO oxidation.<sup>[18,26]</sup>

Germanium has been demonstrated to be a promising promoter for Pt. In the past, it has been used to boost the performance of Pt catalysts in a variety of reactions such as naphtha reforming<sup>[27–29]</sup> and alkane dehydrogenation.<sup>[30–36]</sup> Similarly, Pt–Ge catalysts were shown to exhibit an enhanced performance for the methanol electrooxidation.<sup>[37]</sup> In this study a reduced CO poisoning over Pt was detected, but it was proposed to arise from the surface functionalization with citric/nitric acids. Later studies showed that the adsorption properties of CO on a Pt/C surface change upon the addition of GeO<sub>2</sub>.<sup>[38]</sup> In addition, Pt–Ge/C catalyst presents an easier CO oxidation as well as a higher electrocatalytic activity in methanol oxidation, probably due to geometric effects.<sup>[39]</sup> Ge was also tested as a promoter in Pt/SiO<sub>2</sub> for the CO oxidation. The experiments show that the activity and selectivity were enhanced by 2 times and 5 times, respectively due to the ability of Ge to activate O<sub>2</sub>.<sup>[40]</sup> In this vein, a theoretical work suggested that M<sub>2</sub>Ge<sub>12</sub> compounds (M = Cr, Mn, Fe, Co and Ni) are promising catalysts for low-temperature CO oxidation, where the reaction barrier has a linear relation with the O<sub>2</sub> binding strength on the clusters.<sup>[41]</sup> Previously, we showed that alloying small-size Pt clusters with Ge significantly diminishes the CO poisoning, as the interaction between Pt and CO is weakened upon alloying.<sup>[42,43]</sup> Furthermore, the composition in such bimetallic catalysts might have a great influence in their physico-chemical properties.<sup>[33,43]</sup>

Building from these results, in this work, we combine first-principles calculations with microkinetic modeling to determine the role of Ge in the CO oxidation by Pt–Ge bimetallic clusters supported on MgO(100). Our aim is to verify whether the reduction of CO poisoning found thereof translates into an improvement of the catalytic activity and what Pt–Ge composition provides the best performance for this reaction. The influence of the Ge concentration on the CO and O<sub>2</sub> adsorption properties and the catalytic behavior is thoroughly studied. A concise mechanistic study of CO oxidation is carried out by exploring and unraveling the kinetic details of the two main pathways, *viz.*, Langmuir–Hinshelwood (LH) and Eley–Rideal (ER), along with the electronic reasons behind the observed behavior. In this way, we demonstrate that it is possible to change the bistable kinetics towards a reactive and CO poisoning-free region by adjusting the promoter content on Pt–Ge catalysts.

## Computational Methods

### DFT Calculations

Density functional (DFT) calculations were done using the Perdew–Burke–Ernzerhof (PBE) exchange–correlation functional<sup>[44,45]</sup> and the projector augmented wave (PAW) pseudopotentials<sup>[46]</sup> as implemented in the Vienna ab-initio simulation package (VASP).<sup>[47–50]</sup> The plane wave cutoff energy was set at 450 eV and a convergence criteria of 10<sup>–6</sup> eV and 10<sup>–5</sup> eV was employed for the electronic energy and the geometry relaxation, respectively. In addition, we included Grimme's semiempirical DFT-D3 corrections<sup>[51]</sup> to account

for the dispersion interactions. The Brillouin zone was sampled using a  $1 \times 1 \times 1$  Monkhorst-Pack mesh centered at the  $\Gamma$  point.

The Pt–Ge clusters were supported on a MgO(100)  $4 \times 4 \times 3$  slab composed of 6 layers. The bottom two layers of the support were fixed at the positions of the bulk, while the top four layers were allowed to relax. Periodic images in  $z$  direction were separated with ca. 15 Å of vacuum. The same ensemble of  $\text{Pt}_4\text{Ge}_n/\text{MgO}(100)$  ( $n = 0-4$ ) isomers as reported in ref. [33] were used in this work. CO and  $\text{O}_2$  molecules were adsorbed in all possible orientations on these catalysts with the global optimization toolkit PGOPT.<sup>[52]</sup> PGOPT uses the bond length distribution algorithm (BLDA) to generate a pool of initial adsorption geometries. Transition states were located with the climbing image nudged elastic band (CI-NEB) method.<sup>[53,54]</sup> It must be mentioned that in a few cases a preliminary CI-NEB run with low convergence criteria was performed followed by a tighter convergence calculation using the improved dimer method.<sup>[55,56]</sup> Due to the high computational cost, we only included the 2 uppermost MgO layers and the criteria for the forces was set to  $0.05 \text{ eV \AA}^{-1}$ . The nature of the transition states was confirmed with a vibrational analysis, making sure that they only displayed a single imaginary frequency along the reaction coordinate.

The binding energy of CO and  $\text{O}_2$  was calculated with Equation (1):

$$BE[X] = E(\text{Pt}_4\text{Ge}_n(X)/\text{MgO}) - E(\text{Pt}_4\text{Ge}_n/\text{MgO}) - E(X) \quad (1)$$

where  $E(\text{Pt}_4\text{Ge}_n(X)/\text{MgO})$  is the energy of  $\text{Pt}_4\text{Ge}_n$  cluster ( $n = 0-4$ ) supported on MgO with one  $X$  molecule adsorbed,  $E(\text{Pt}_4\text{Ge}_n/\text{MgO})$  is the energy of the bare  $\text{Pt}_4\text{Ge}_n$  cluster supported on MgO and  $E(X)$  is the energy of the isolated  $X$  molecule in gas phase ( $X = \text{CO}, \text{O}_2$ ). The same reference  $E(\text{O}_2)$  is used for  $\text{O}_2$  and  $2 \times \text{O}$  adsorption.

### Microkinetic modeling

Microkinetic modeling is a powerful technique used to analyze complex catalytic networks in terms of elementary reactions.<sup>[57,58]</sup> Herein, no assumptions are made, a priori, about the rate-determining steps or surface coverage of intermediate species. In this fashion, measured or calculated parameters such as activation and reaction free energies can be incorporated into the rate expressions. Regarding the reactor model, a continuous stirred-tank reactor (CSTR) was chosen to solve the rate equations. In this way, the time evolution of the coverages and pressures of each species can be obtained. The details of the reaction network are given in the Supporting Information. The simulations were run for at least  $10^7 \text{ s}$  to ensure convergence to a dynamic steady-state. An important point should be made here mentioning that the achieved dynamic steady-state differs significantly from the thermodynamic equilibrium as there is a continuous gas inlet and outlet which brings the system out of equilibrium in a dynamic way.<sup>[59]</sup> The total flow rate was set to  $8 \text{ ml min}^{-1}$  with different  $\text{O}_2$ :CO compositions. 100 mg of catalyst was used with a Pt loading of 0.35 wt%, based on previous experimental work.<sup>[28,60]</sup> Pt and Ge sites were treated equivalently in the spirit of the mean-field approach.

For surface reactions, the rate constants for the forward elementary reactions were determined by the transition state theory (TST) [Equation (2)]:

$$k_{f,i} = \frac{k_B T}{h} e^{-\frac{\Delta G_i^\ddagger}{k_B T}} \quad (2)$$

where  $k_B$  is Boltzmann's constant,  $h$  is Planck's constant and  $T$  is the temperature.  $\Delta G_i^\ddagger$  is the Gibbs free activation energy of the  $i^{\text{th}}$

elementary reaction. The equilibrium constants  $K_i$  were calculated as follows [Equation (3)]:

$$K_i = e^{-\frac{\Delta G_i}{k_B T}} \quad (3)$$

where  $\Delta G_i$  is the Gibbs free energy of the  $i^{\text{th}}$  elementary reaction. To ensure thermodynamic consistency the reverse rate constants were calculated using the corresponding forward rate constant and equilibrium constant [Equation (4)]:

$$K_i = \frac{k_{f,i}}{k_{r,i}} \quad (4)$$

The ideal gas limit was considered to compute the changes in free energy [Equation (5)]:

$$\Delta G = \Delta(E_{\text{elec}} + \text{ZPE}) - T\Delta S \quad (5)$$

where  $E_{\text{elec}}$  is the DFT computed electronic energy, ZPE is the zero-point energy and  $\Delta S$  is the change in entropy. The zero-point energy was calculated from the vibrational analysis. For adsorbed species only the harmonic vibrational contribution to entropy was considered, while translational and rotational entropies were also included for gas phase species.

For non-activated unimolecular adsorption, the rate of adsorption is determined by the kinetic theory of gases [Equation (6)]:<sup>[61]</sup>

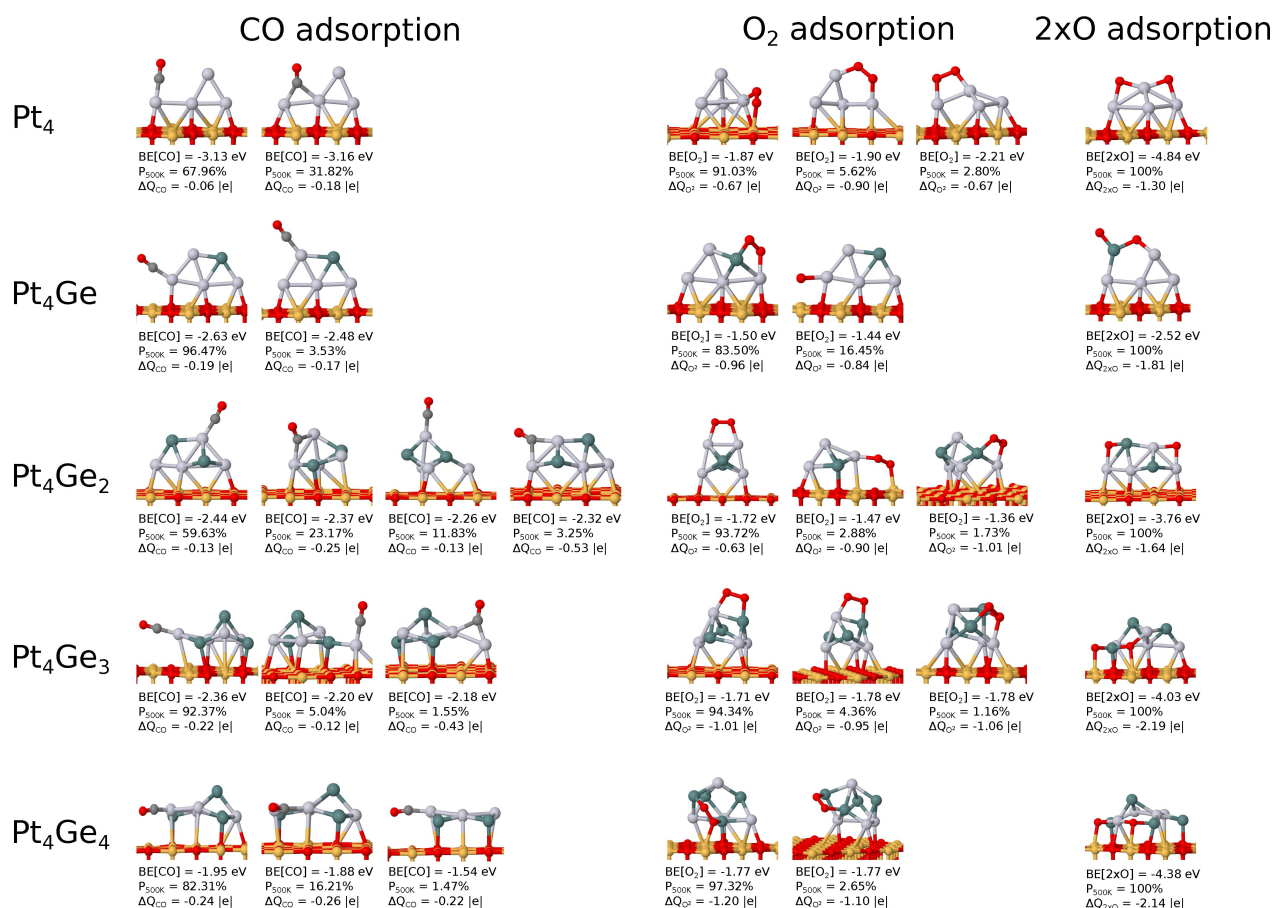
$$k_{\text{ads}} = \frac{PA}{\sqrt{2\pi mk_B T}} S \quad (6)$$

where  $P$  is the partial pressure of the adsorbate,  $m$  is the mass of the adsorbate and  $A$  is the area of the adsorption site. The adsorption site area was set equal to the area of the adjacent triangle. For  $\text{O}_2$  chemisorption (bridge site), the area was approximated as the area of a triangle of sides Pt–Pt–Pt. For CO chemisorption (top site), the area was set to one-third of the area for  $\text{O}_2$ .  $S$  is the sticking coefficient and was set to 1 for all species, based on their favorable adsorption energies.

## Results and Discussion

### CO and $\text{O}_2$ Adsorption on $\text{Pt}_4\text{Ge}_n/\text{MgO}(100)$ ( $n = 0-4$ )

Before moving to the discussion of the CO reaction mechanisms we first study the adsorption of the reactants, viz., CO and  $\text{O}_2$ , on  $\text{Pt}_4\text{Ge}_n/\text{MgO}$  with  $n = 0-4$ . These surface-mounted nano-clusters have been recently characterized and are predicted to be synthetically accessible under Ge-rich conditions.<sup>[33]</sup> For each composition, various conformers would coexist since they are energetically competitive, except for  $\text{Pt}_4\text{Ge}$  in which a single isomer, the global minimum (GM), dominates the population. Note that, while all the bimetallic clusters have a closed-shell electronic structure,  $\text{Pt}_4/\text{MgO}$  possesses unpaired electrons in most cases.<sup>[33]</sup> Because isomers can exhibit different reactivity, CO and  $\text{O}_2$  were adsorbed on the most populated conformers of each  $\text{Pt}_4\text{Ge}_n$  composition. The resulting most stable adsorption geometries along with the CO and  $\text{O}_2$  binding energies (BE), Boltzmann populations of each structure at 500 K, and Bader effective charges are summarized in Figure 1. The first



**Figure 1.** Most stable CO, O<sub>2</sub> and 2xO adsorption geometries on Pt<sub>4</sub>Ge<sub>*n*</sub>/MgO(100) clusters along with the ligand binding energies (*BE*) in units of eV, Boltzmann populations at 500 K (*P*<sub>500K</sub>) and charge transfer to ligand ( $\Delta Q$ ) in values of |e|. Only isomers with *P*<sub>500K</sub> > 1.00% are shown. Pt, Ge, Mg, O and C are depicted in gray, dark green, orange, red and dark gray, respectively.

thing to note is that after adsorption the size of the ensemble and the relative stability of the isomers changes, showing the dynamic effect of adsorption.<sup>[15]</sup> CO adsorption has a profound effect on the cluster population. For instance, in Pt<sub>4</sub>Ge<sub>3</sub>, isomer IV in ref [33] (population at 500 K of 8.3%) becomes predominant upon CO adsorption, while the most stable bare clusters become negligible. In general, CO prefers to adsorb on Pt through the C end in an atop configuration ( $\angle\text{Pt-C-O } 180^\circ$ ), with a few exceptions, where it binds in a bridge fashion. CO adsorption has little impact on the cluster geometry, and CO is never found adsorbed on Ge which indicates the much higher preference towards Pt. O<sub>2</sub> is found to adsorb either dissociatively (2xO\*) or non-dissociatively (O<sub>2</sub>\*). Contrary to CO, oxygen can bind to both Pt and Ge, the most common active site being a bimetallic Pt–Ge ensemble. In the dissociative adsorption atomic O binds in a M–O–M bridge configuration, while in the non-dissociative adsorption it usually adopts a M–O–O–M two-member bridge site. It should be noted that in many isomers O\* adsorbs at the cluster-support interface bonded to a Mg atom, which is in stark contrast to CO. Moreover, a significant distortion is observed upon adsorption due to the oxophilicity of Ge. In some adducts (e.g., Pt<sub>4</sub>Ge<sub>2</sub>O isomer I) O\* opens the

Pt–Ge polar bonds deforming the structure. This effect may become more significant at strong oxidizing conditions.

The Sabatier principle states that the binding strength between reactants and catalysts should not be very strong to avoid overbinding and poisoning effects, nor very weak to give enough time for diffusion and collision. In addition, the products should interact weakly with the surface to guarantee a facile desorption. In Figure 1, pure Pt<sub>4</sub>CO presents the largest CO binding energy (isomer I, *BE*[CO] = −3.13 eV), suggestive of a tendency to undergo CO poisoning. Nonetheless, adding a single Ge atom in Pt<sub>4</sub>GeCO, the *BE*[CO] is reduced to −2.63 eV; and as the Ge concentration increases a gradual drop in the *BE*[CO] is observed. This trend is in line with previous observations in gas-phase and graphene-supported Pt–Ge nanoclusters.<sup>[42,43]</sup> The bonding between Pt and CO can be understood in terms of the Blyholder model.<sup>[62]</sup> Briefly, it describes the interaction with an electron transfer from the 5σ of CO to the empty Pt d orbitals (donation step), and an electron transfer from the Pt d orbitals to the empty 2π\* of CO (back-donation step). According to the Bader analysis (Figure 1) and projected density of states (PDOS, Figure S3), there is a net charge transfer from the cluster to CO 2π\*. As a result, the C–O

bond is weakened and elongated slightly from 1.14 Å in gas-phase to ~1.17 Å.

As stated above, for O<sub>2</sub> both the molecular and dissociative adsorptions are competitive. Comparing the binding energies to oxygen it is evident that the latter is more stable than the former; as explained in the next section, O<sub>2</sub> dissociation is an exothermic process. Overall CO is more strongly adsorbed than molecular oxygen. The same is true for atomic oxygen in monometallic and Pt<sub>4</sub>Ge clusters. However, as the Ge concentration increases, the adsorption of 2×O\* becomes stronger than that of CO\* by more than 2 eV in Pt<sub>4</sub>Ge<sub>3</sub> and Pt<sub>4</sub>Ge<sub>4</sub>, meaning that O\* may become the predominant species on the catalyst surface. Regardless of the adsorption mode, O atoms carry a negative charge owing to an electron transfer from the cluster to the empty 2π\* orbitals. As expected, dissociated O\* has a significantly larger charge than molecular O<sub>2</sub>\* since a single atom is interacting with two Pt or Ge atoms. Depending on the degree of activation (O–O elongation) molecular O<sub>2</sub>\* can change its electronic state. In gas-phase the diatomic distance is 1.21 Å (bond order = 2). Upon adsorption, the internal bond weakens (bond order reduction) and elongates. If one electron is transferred to oxygen, the superoxo (O<sub>2</sub><sup>-</sup>) state is obtained with a O–O bond distance of ~1.33 Å (bond order ~1.5). O<sub>2</sub>\* can still accept another electron and achieve the peroxo (O<sub>2</sub><sup>2-</sup>) state with a O–O bond distance of ~1.47 Å (bond order ~1.0). Our calculations suggest a large oxygen activation for the Pt–Ge nanoclusters. The internal bond is increased up to 1.47–1.50 Å in the bidentate bridge configuration, indicating the presence of a peroxo state; while in the atop configuration it is only activated until the superoxo state with a bond length of 1.37 Å. We believe that such a high activation translates into low oxygen dissociation barriers.

### O<sub>2</sub> Dissociation on Pt<sub>4</sub>Ge<sub>n</sub>/MgO(100) (n = 0–4)

To investigate the O<sub>2</sub> molecular and dissociative adsorption mechanisms, we have computed the O<sub>2</sub> dissociation barriers. The energy profiles for the different Pt<sub>4</sub>Ge<sub>n</sub> clusters are collected in Figure 2. Here, we only show the results for the pathways with the most stable Pt<sub>4</sub>Ge<sub>n</sub>2O final state (FS) adducts for each cluster; the rest of the isomers are depicted in Figures S4–S7 of the SI.

After O<sub>2</sub> chemisorption, the cluster activates the O–O bond by transferring electrons to the antibonding 2π\* orbitals. For the bridge adsorption mode O<sub>2</sub> is activated in the peroxo state where the IS structure is close to the transition state (TS). The cluster continues to stretch the bond until finally the O–O scission is achieved. In all cases the O–O bond is already cleaved in the TS, even for the atop adsorption (see Figure S4). Our calculations reveal that the O–O bond expands from ~1.47 Å to a value in the range of 1.90–2.15 Å in the TS. Here the charges of the O atoms are maximal; each atom carries 0.6–0.8|e| comparing to ~0.4|e| in the IS. The values of these charges are similar for adsorption on Pt–Pt and Pt–Ge of M–O–O–M bridge sites. For the numerical values of the oxygen Bader charges and O–O distances the reader is referred to

Table S1. Afterwards, the O atoms separate and occupy a bridge site in the vicinity. This process is found to be exothermic, the greatest energy release occurring when the strong Ge–O bonds distort the cluster structure (e.g., Pt<sub>4</sub>Ge<sub>4</sub> isomer II, Figure S7). Overall, the activation barriers are small and most cases lie in the range of 0.4–0.6 eV, although there are some exceptions where the barrier increases to 0.75 eV, as for the Pt<sub>4</sub>Ge<sub>3</sub> isomer I (see Figure S6). Our results indicate that at room temperature most of the adsorbed oxygen will be in the form of atomic O\*.

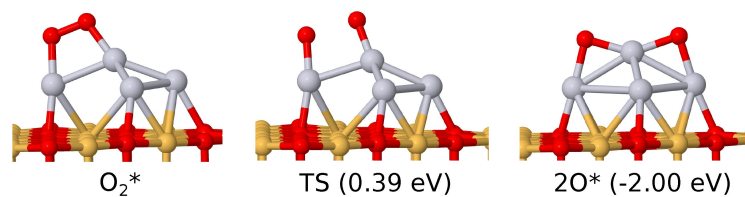
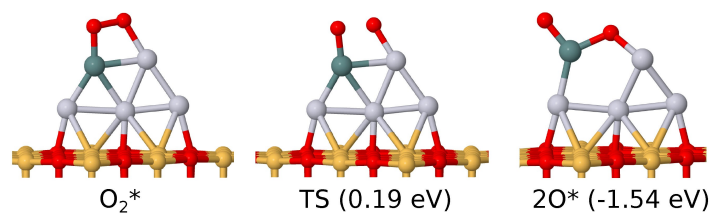
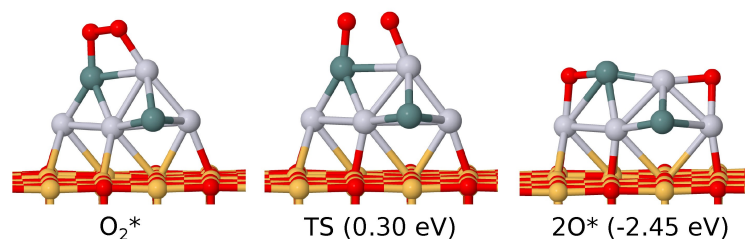
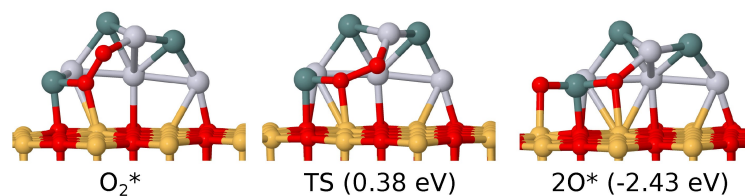
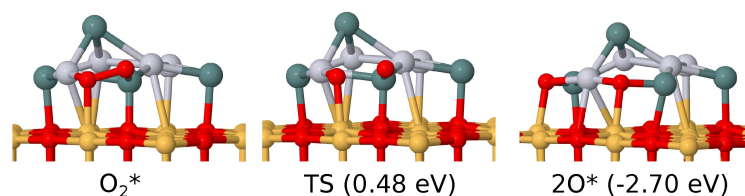
### CO Oxidation on Pt<sub>4</sub>Ge<sub>n</sub>/MgO(100) (n = 0–4)

Here, we characterize the potential energy surface of CO oxidation on Pt<sub>4</sub>Ge<sub>n</sub> clusters (n = 0–4) supported on MgO(100). On Pt-based catalysts, CO oxidation usually proceeds via the bimolecular LH or ER mechanisms. In LH prior to the reaction both CO and O<sub>2</sub> are adsorbed on the surface, that is, co-adsorption of the reactants is required. Then the reactants diffuse to adjacent sites and react with each other to form CO<sub>2</sub>, which is released to the environment. Besides, the ER mechanism only involves O<sub>2</sub> adsorption followed by a CO attack in the gas-phase. As has been confirmed in the previous section, all clusters present small oxygen dissociation barriers. Hence, CO will be able to react not only with O<sub>2</sub>\* but also with O\*; in other words, both O<sub>2</sub> and O are available as oxygen sources. We must highlight that all the possible reaction mechanisms were characterized by considering two different structures for each Pt<sub>4</sub>Ge<sub>n</sub>/MgO, the isomers with the most stable IS and FS in the O<sub>2</sub> dissociation shown in Figures 2 and S4–S7. Pt<sub>4</sub>Ge<sub>n</sub>O<sub>2</sub> was used to study the routes involving CO + O<sub>2</sub>, while Pt<sub>4</sub>Ge<sub>n</sub>2O was employed for the CO + 2×O paths. Since the starting point involves two oxygen atoms in either case, two steps will be needed to complete the catalytic cycle, *i.e.*, the formation of two CO<sub>2</sub> molecules.

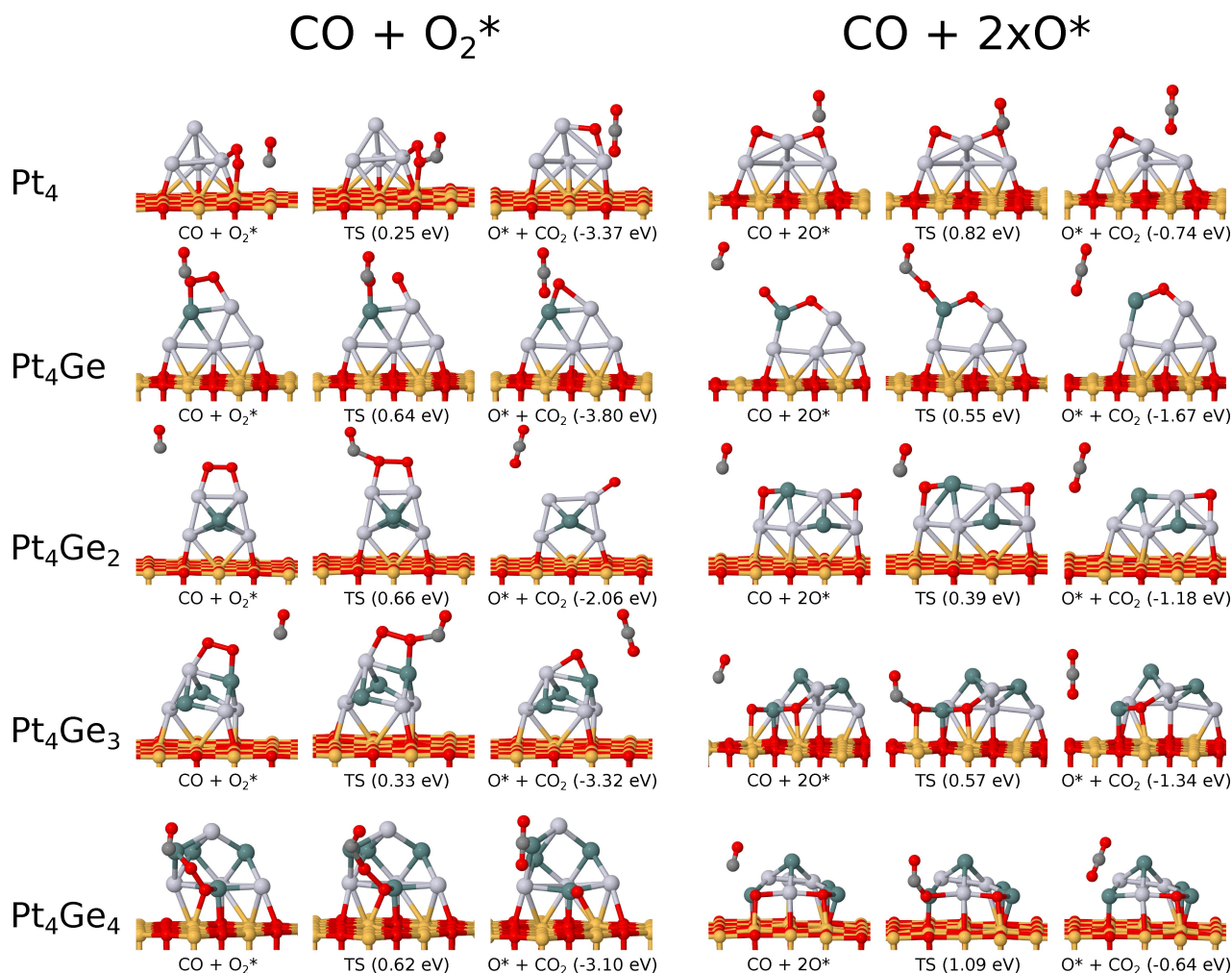
### Eley-Rideal mechanism

The ER mechanism takes place when CO in the gas-phase attacks O<sub>2</sub>\* or predissociated O\* on the surface. Given the small size of the nanoclusters explored in this work, the variety of co-adsorption configurations is limited, and hence the ER mechanism may actually be a competitive mechanism within this size regime. Figure 3 shows the reaction pathways for both CO(g) + O<sub>2</sub>\* and CO(g) + 2×O\* for all Pt<sub>4</sub>Ge<sub>n</sub> considered in this section. For the sake of clarity, we only show the results for the first CO<sub>2</sub> formation in both routes; the second CO<sub>2</sub> formation is depicted in Figure S8.

As depicted in Figure 3, CO(g) attack to O<sub>2</sub>\* always occurs from the C end. As CO approaches the molecules, the O–O bond starts activating until the TS is reached. It should be remarked that the TS is effectively achieved only under certain angles of attack (∠O–C–O\* between 120° and 125°). In the TS, CO(g) is found at a distance of ~1.70 Å from O<sub>2</sub>. The CO internal bond is almost unchanged while O–O is activated to ~1.68 Å. After the energy barrier is successfully overcome, CO<sub>2</sub> is directly

a) Pt<sub>4</sub>b) Pt<sub>4</sub>Gec) Pt<sub>4</sub>Ge<sub>2</sub>d) Pt<sub>4</sub>Ge<sub>3</sub>e) Pt<sub>4</sub>Ge<sub>4</sub>

**Figure 2.** Initial state (IS), transition state (TS) and final state (FS) geometries for the O<sub>2</sub> dissociation pathway on the Pt<sub>4</sub>Ge<sub>*n*</sub>/MgO(100) isomers with the most stable final state (FS). Activation and reaction energies (in eV) are given relative to the corresponding IS. Pt, Ge, Mg, O and C are depicted in gray, dark green, orange, red and dark gray, respectively.



**Figure 3.** Most stable reaction pathways for the Eley-Rideal (ER) mechanism for both molecular and dissociated oxygen reactions on Pt<sub>4</sub>Ge<sub>n</sub>/MgO(100). Initial state (IS), transition state (TS) and final state (FS) geometries are shown together with the relative activation and reaction energies (in eV). Only the first step is given. Pt, Ge, Mg, O and C are depicted in gray, dark green, orange, red and dark gray, respectively.

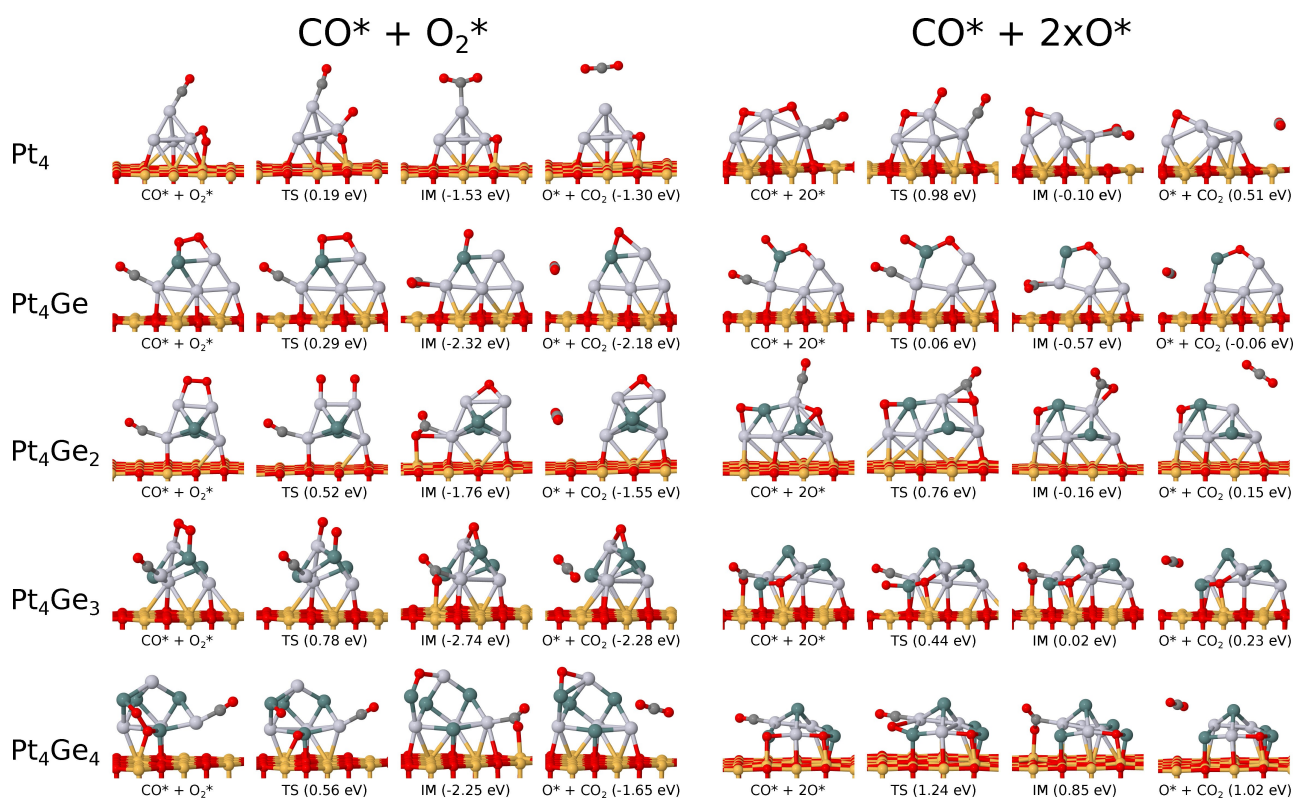
ejected from the surface without the formation of an intermediate state (IM). The corresponding energy barriers for this route are rather small; Pt<sub>4</sub> has a barrier of 0.25 eV, whereas in the nanoalloys it lies in the range of 0.30–0.60 eV. In contrast, in the most likely route, the CO(g) + 2O\* route (Figure 3), the barriers on Pt<sub>4</sub> increase up to 0.80 eV. Pt<sub>4</sub>Ge (second step) and Pt<sub>4</sub>Ge<sub>4</sub> (first step) also show high barriers of 1.0 eV, whereas in the clusters with medium Ge content they remain below 0.60 eV. All steps are found to be exothermic. These results suggest that Pt<sub>4</sub>Ge<sub>2</sub> and Pt<sub>4</sub>Ge<sub>3</sub> are the most active compositions. Nonetheless, as long as there is predissociated O\*, *i.e.*, CO poisoning is avoided, the reaction will move forward in all cases and thus, the ER channel will proceed similarly regardless of the Pt–Ge composition.

### Langmuir-Hinshelwood mechanism

In Figure 4 we display the reaction pathways through the first step of the LH channel for CO oxidation by both molecular and

atomic oxygen on Pt<sub>4</sub>Ge<sub>n</sub> clusters; the second step is depicted in Figure S9.

To search for the most suitable starting point a screening of the most stable co-adsorption complexes was done by adsorbing CO with multiple orientations on adjacent sites to O<sub>2</sub>\* or O\* in Pt<sub>4</sub>Ge<sub>n</sub>O<sub>2</sub> and Pt<sub>4</sub>Ge<sub>n</sub>2O. We begin with the CO\* + O<sub>2</sub>\* reaction (Figure 4). As the reactants approach each other the O–O bond is elongated and the reaction proceeds via the first TS to an intermediate (IM). It should be mentioned that in many cases<sup>[63]</sup> this first step involves the formation of a OCOO\* species in the IM. However, our calculations for Pt–Ge reveal that the cleavage of O–O and formation of C–O occurs simultaneously in the first step, so CO<sub>2</sub>\* is already formed in the IM. Apart from Pt<sub>4</sub>Ge, oxygen is already dissociated in the TS. The charge of the reacting O increases up to ~1.0 |e|. In the IM complex CO<sub>2</sub> is usually bonded from the C atom to Pt with a Pt–C distance of ~2.05 Å and has an internal angle of ~130°. When the reaction occurs close to the interface, CO<sub>2</sub> also binds to a substrate Mg atom via one of the oxygens. Subsequently, there is a barrierless



**Figure 4.** Most stable reaction pathways for the Langmuir-Hinshelwood (LH) mechanism for both molecular and dissociated oxygen reactions on  $\text{Pt}_4\text{Ge}_n/\text{MgO}(100)$ . Initial state (IS), transition state (TS), intermediate state (IM) and final state (FS) geometries are shown together with the relative activation and reaction energies (in eV). Only the first step is given. Pt, Ge, Mg, O and C are depicted in gray, dark green, orange, red and dark gray, respectively.

release of the formed  $\text{CO}_2$  molecule to the environment, leaving an  $\text{O}^*$  on the surface to further oxidize another CO molecule.

Pure  $\text{Pt}_4$  presents small barriers for the first (0.19 eV) as well as the second step (0.50 eV). Regarding the bimetallic counterparts, the first step for  $\text{Pt}_4\text{Ge}$  and  $\text{Pt}_4\text{Ge}_2$  are also rather small (0.29 and 0.52 eV), whereas the second barrier becomes significantly large ( $> 1.5$  eV), while  $\text{O}^*$  has to migrate from one very stable bridge position to the adjacent one. In  $\text{Pt}_4\text{Ge}_3$  the first  $\text{CO}_2$  formation shows a barrier of 0.78 eV, while in  $\text{Pt}_4\text{Ge}_4$  both barriers are around 0.55 eV. The first step is always more exothermic than the second as the O–O scission releases extra energy.  $\text{CO}_2$  desorption is found to be endothermic with a  $BE$  of around 0.10 eV and 0.30 eV in the first and second steps, respectively.

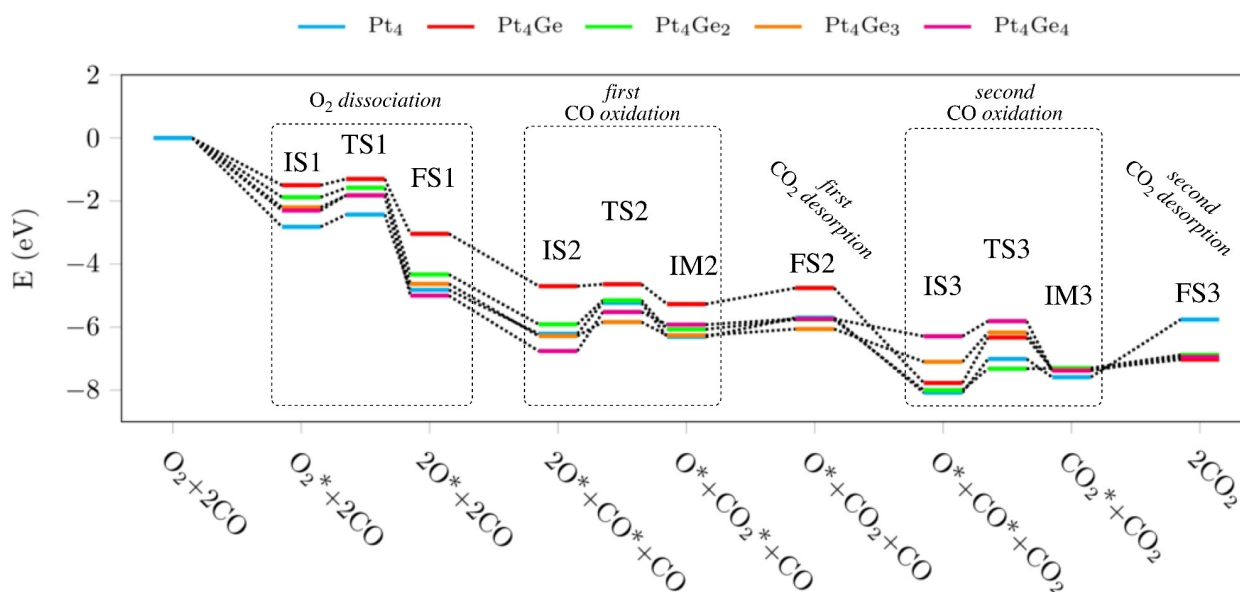
The most likely scenario, however, is the one where  $\text{CO}^*$  reacts with predissociated  $\text{O}^*$  ( $\text{CO}^* + 2\times\text{O}^*$ ), as shown in Figure 4. A detailed inspection reveals that in the TS of this channel  $\text{CO}^*$  and  $\text{O}^*$  are closer than in  $\text{CO}^* + \text{O}_2^*$ . The reactants are bonded on top of neighbouring atoms and oriented parallel to each other, forming a cyclic Pt–C–O–Pt(Ge) structure. The length of the “to be formed” C–O bond lies between 2.75 and 2.90 Å. Such TS structure has also been observed previously in the CO oxidation by small Pd clusters.<sup>[64]</sup> Contrary to the previous pathway,  $\text{Pt}_4$  has to overcome now larger barriers of  $\sim 1.0$  eV in both steps. Surprisingly, on  $\text{Pt}_4\text{Ge}$  the first step proceeds almost barrierless, although the second oxidation is highly unfavorable. On the contrary,  $\text{Pt}_4\text{Ge}_2$  and  $\text{Pt}_4\text{Ge}_3$  show

the best balance between the two steps with low barriers. Moreover, in this route the first IM formation is significantly less exothermic since the reaction starts from a very stable complex. Except for  $\text{Pt}_4\text{Ge}_4$  (0.85 eV), the first IM formation is still slightly exothermic, while the second step is predicted to be exothermic only on  $\text{Pt}_4\text{Ge}_3$  and  $\text{Pt}_4\text{Ge}_4$  (see Figure S9). Another observation of critical importance is the fact that  $\text{CO}_2$  removal becomes highly endothermic in  $\text{Pt}_4$ , especially in the second step, as opposed to the alloys. This suggests a tendency to get poisoned by  $\text{CO}_2$ .

It is extremely important to consider the statistical weights of each reaction channel. Indeed, the most likely reaction mechanism for LH will be connected to atomic oxygen. In this case bimetallic clusters present lower energy barriers as compared to pure Pt, as well as the inhibition of the CO and  $\text{CO}_2$  poisoning effects.

For the sake of comparison, the complete potential energy diagram of the LH mechanism with predissociated oxygen on all the  $\text{Pt}_4\text{Ge}_n$  nanoclusters is illustrated in Figure 5 (the reader is referred to Figure S10 for the results of the other cases). It can be observed that the LH mechanism lies lower in energy than ER since it includes the adsorption of both  $\text{O}_2$  and CO, while ER is missing the adsorption of CO. In the ER mechanism for both  $\text{CO}(\text{g}) + \text{O}_2^*$  and  $\text{CO}(\text{g}) + 2\times\text{O}^*$  (Figure S10), all steps are exothermic (downhill) for all clusters, the  $\text{CO}_2$  release being the most stable state. Consequently, this state will act as a thermodynamic sink shifting the reaction towards  $\text{CO}_2$  forma-





**Figure 5.** Potential energy diagram of the most stable  $\text{CO}^* + 2\text{O}^*$  routes (LH mechanism) on  $\text{Pt}_4\text{Ge}_n/\text{MgO}(100)$ . Energies (in eV) are given relative to the clean  $\text{Pt}_4\text{Ge}_n/\text{MgO}(100)$  surface.

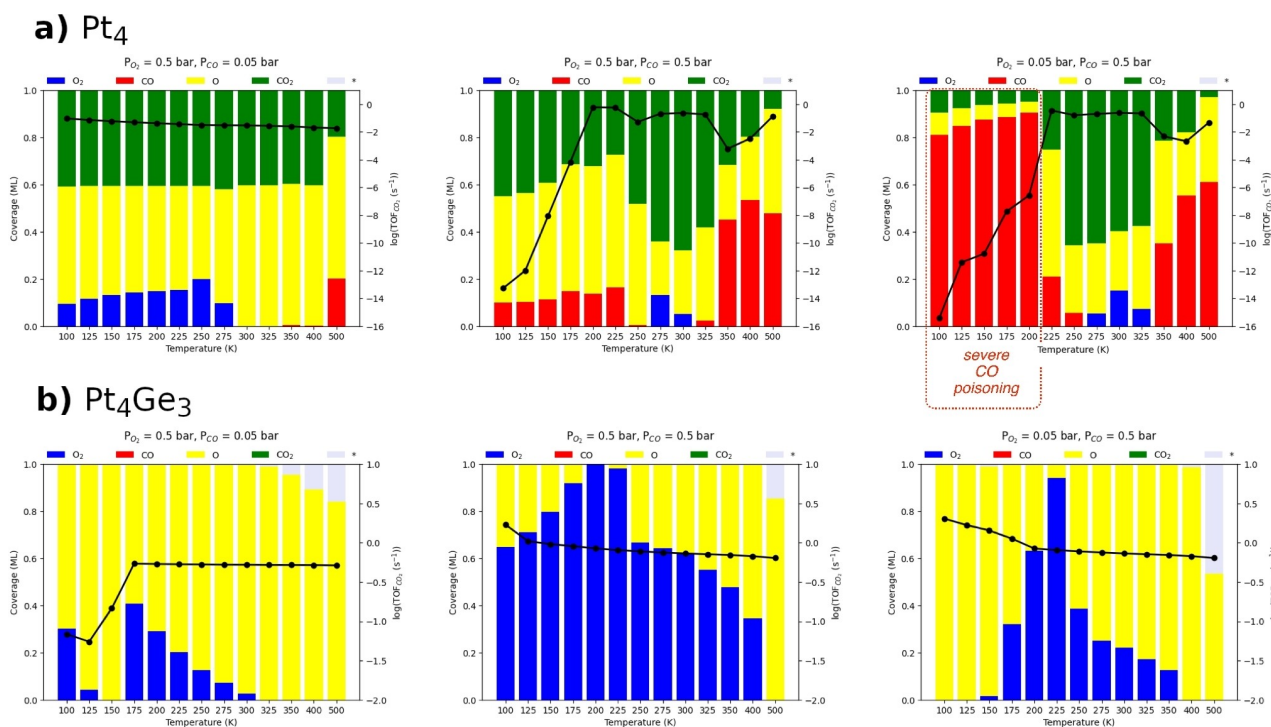
tion, regardless of the Ge concentration. Notwithstanding, as shown in Figure 5, in the LH mechanisms  $\text{O}^* + \text{CO}^*$  ( $\text{IS}_3$ ,  $\text{O}^* + \text{CO}^*$ ) becomes the most stable state for  $\text{Pt}_4$ ,  $\text{Pt}_4\text{Ge}$  and  $\text{Pt}_4\text{Ge}_2$  due to the endothermic nature of the last two steps (second oxidation + desorption). This indicates that  $\text{IS}_3$  will act as a drain and shift the reaction towards  $\text{CO}^* + \text{O}^*$  formation. In this state  $\text{O}^*$  could still react with  $\text{CO}$  in gas-phase via ER mechanism, but part of the surface will be poisoned by  $\text{CO}$ . The case of  $\text{Pt}_4$  is more dramatic since  $\text{CO}_2$  desorption is highly endothermic, which means that part of the formed  $\text{CO}_2$  will remain on the surface, *i.e.*,  $\text{Pt}_4$  will be poisoned by  $\text{CO}_2$  as well as  $\text{CO}$ .

Contrary to pure Pt, the formation of the second  $\text{CO}_2^*$  ( $\text{IM}_3$ ) becomes the lowest stationary point when increasing the Ge content, *i.e.*, in  $\text{Pt}_4\text{Ge}_3$  and  $\text{Pt}_4\text{Ge}_4$ . Because  $\text{CO}_2$  desorption is slightly endothermic in both cases,  $\text{CO}_2$  will easily detach from the cluster. On  $\text{Pt}_4\text{Ge}_3$  the oxidation steps are somewhat reversible, while on  $\text{Pt}_4\text{Ge}_4$  the first step is very endothermic. From these observations it can be concluded that in either case the surface will be covered by  $\text{CO}^*$ ,  $\text{O}^*$  and  $\text{CO}_2^*$ , but in  $\text{Pt}_4\text{Ge}_3$  once  $\text{CO}_2(\text{g})$  is formed, the reaction will evolve more easily towards the final product. Furthermore, these two clusters present a higher  $BE$  for  $\text{O}_2$  than  $\text{CO}$ , so  $\text{O}^*$  may become the predominant species together with  $\text{CO}_2^*$ , limiting  $\text{CO}$  poisoning. The  $\text{O}^*$  in excess should not be considered detrimental since it can still react with  $\text{CO}(\text{g})$  via the ER channel. In reality,  $\text{CO}$  is being consumed simultaneously in each one of these routes; in other words, the different pathways are not independent of each other and the reaction should be thought of as a dynamic phenomenon where the coverage of the intermediates obtained through one route may affect the efficiency of the other. Such a relationship cannot be captured with the results obtained so far. In this respect, microkinetic modeling can be a very useful tool to help elucidate the time evolution of complex reaction networks.

### Microkinetic simulations

To shed light on the effect of the alloy concentration over the time evolution of  $\text{CO}_2$  production, the DFT calculations were complemented with an 8-step microkinetic model, following the approach of previous work.<sup>[65,66]</sup> Microkinetic modeling allows us to make a more direct comparison between DFT calculations and experimental variables such as partial pressures and temperatures. Importantly, the coverage effects are also considered, offering a thorough assessment of the catalyst poisoning tendency. In the simulations reported herein, the temperature was varied from 100 to 500 K and the  $\text{O}_2/\text{CO}$  ratio of the feed was altered to simulate different environments (*viz.*,  $P_{\text{O}_2}=0.5$  bar,  $P_{\text{CO}}=0.5$  bar;  $P_{\text{O}_2}=0.5$  bar,  $P_{\text{CO}}=0.05$  bar;  $P_{\text{O}_2}=0.05$  bar,  $P_{\text{CO}}=0.5$  bar). It should be stressed that no a priori assumptions were made about the rate-determining step or steady-state surface coverages. The elementary steps and reactor model are described in greater detail in the SI.

In Figure 6 we compare the steady-state coverages and the  $\text{CO}_2$  turnover frequency ( $\text{TOF}_{\text{CO}_2}$ ) as a function of temperature and  $\text{CO}$  and  $\text{O}_2$  partial pressures of  $\text{Pt}_4$  (pure Pt) with the best alloy candidate, *i.e.*,  $\text{Pt}_4\text{Ge}_3$ ; the rest are shown in Figure S12, along with the discussion. A quick glance reveals that  $\text{Pt}_4$  exhibits the well-reported bistable kinetics in agreement with previous studies.<sup>[6,67,68]</sup> Different regimes of slow (severe  $\text{CO}$  poisoning; Figure 6a, right) and fast ( $\text{CO}$  poisoning-free; Figure 6a, left) kinetics can be distinguished based on the surface coverages. High  $\text{O}_2$  pressures are the best conditions to avoid  $\text{CO}$  poisoning and increase the  $\text{CO}_2$  formation rate. At equal  $\text{CO}$  and  $\text{O}_2$  pressures (Figure 6a, middle),  $\text{O}^*$  ( $\theta_{\text{O}} \sim 0.5$  ML) and  $\text{CO}_2^*$  ( $\theta_{\text{CO}_2} \sim 0.4$  ML) become the main species on the surface in the range of 100–200 K. Interestingly, despite using a feed with a  $\text{CO}:\text{O}_2$  1:1 composition, the  $\text{CO}$  coverage only amounts to 0.1 ML. In this temperature regime,  $\text{Pt}_4$  is more reactive in the



**Figure 6.** Microkinetic analysis of CO oxidation for a) Pt<sub>4</sub> and b) Pt<sub>4</sub>Ge<sub>3</sub> clusters supported on MgO(100). Steady-state coverages and CO<sub>2</sub> turnover frequencies (TOF<sub>CO<sub>2</sub></sub>) as a function of temperature with different feed compositions: P<sub>O<sub>2</sub></sub> = 0.5 bar, P<sub>CO</sub> = 0.05 bar (left figure); P<sub>O<sub>2</sub></sub> = 0.5 bar, P<sub>CO</sub> = 0.5 bar (middle figure); P<sub>O<sub>2</sub></sub> = 0.05 bar, P<sub>CO</sub> = 0.5 bar (right figure).

CO\* + O<sub>2</sub>\* route than in the CO\* + O\* one (see Table S3), making the former the predominant path for CO<sub>2</sub> production. The high endothermicity of CO<sub>2</sub> desorption makes CO<sub>2</sub>\* block an important portion of the free active sites. In addition, the O\* produced via CO\* + O<sub>2</sub>\*, remains mostly unreacted owing to the high energy barrier in the CO\* + O\* route (see Table S3), which becomes reflected in the low TOF<sub>CO<sub>2</sub></sub> values. At intermediate temperatures (200–325 K) CO<sub>2</sub>\* becomes the predominant species which may be ascribed to an increment in the CO<sub>2</sub> desorption free energy. The TOF<sub>CO<sub>2</sub></sub> also increases and remains constant at a fairly moderate level (TOF<sub>CO<sub>2</sub></sub> ~ 10<sup>-1</sup> s<sup>-1</sup>). It should be noted that O\* is constantly reacting with CO in gas-phase via ER mechanism. The empty sites left will either become covered with O\* again or CO\* and CO<sub>2</sub>\* (the latter being more likely). Then, from 350 K upwards, a change of mechanism occurs and CO\* + O\* becomes the main route (see Table S3). The barriers for O<sub>2</sub> dissociation and CO\* + O\* reaction become lower than the CO\* + O<sub>2</sub>\* ones. Yet, the difference in binding free energies (BGs) between CO and O<sub>2</sub> becomes more important, and therefore CO poisoning also becomes more pronounced, slowing down the reaction rate.

At high O<sub>2</sub> pressures (Figure 6a, left) a transition can be observed to the CO poisoning-free regime. Higher O<sub>2</sub> pressures seem to be beneficial, regardless of the temperature, as expected. O<sub>2</sub> adsorption becomes more favorable than CO (see Table S3), which then converts to O\* via O<sub>2</sub> dissociation and CO\* + O<sub>2</sub>\*. From 100 K to 275 K the surface composition is 0.4 ML CO<sub>2</sub>\*, ~0.5 ML O\* and ~0.1 ML O<sub>2</sub>\*. At higher temperature, O<sub>2</sub> dissociation becomes more favorable and O\* coverage

increases at the cost of O<sub>2</sub>\* sites. The presence of CO<sub>2</sub>\* indicates that the LH channel is not hindered by CO poisoning. However, CO<sub>2</sub>\* limits the availability of active sites for LH and ER pathways, so the efficiency of CO<sub>2</sub> production is compromised. The TOF<sub>CO<sub>2</sub></sub> decreases linearly with temperature from 10<sup>-1</sup> to 10<sup>-2</sup> s<sup>-1</sup>. We anticipate that CO<sub>2</sub> poisoning can cause a drop in the TOF<sub>CO<sub>2</sub></sub> by two orders of magnitude, although the effect is milder than for CO poisoning.

A CO-rich feed (Figure 6, right) has an extremely detrimental effect on the performance of the monometallic cluster at low temperatures; Pt<sub>4</sub> becomes almost completely poisoned by CO (θ<sub>CO</sub> ~ 0.8 ML). In this regime there is a scarcity of O<sub>2</sub>\* and O\* sites for LH and ER paths and thus the oxidation rate decreases drastically. This behavior can be understood in terms of DFT binding free energies that show the strong adsorption of CO on pure Pt clusters. The TOF<sub>CO<sub>2</sub></sub> increases with temperature until it is stabilized in the 225–325 K window. CO\* nearly disappears but CO<sub>2</sub>\* coverage becomes more pronounced (θ<sub>CO<sub>2</sub></sub> ~ 0.6 ML). Nonetheless, there is still a fraction of O\* and O<sub>2</sub>\* which can oxidize CO at rates similar to P<sub>O<sub>2</sub></sub> > P<sub>CO</sub>. As the temperature increases, CO\* coverage starts increasing at the expense of CO<sub>2</sub>\* and the TOF<sub>CO<sub>2</sub></sub> is slightly reduced.

We must emphasize that CO and CO<sub>2</sub> poisoning are completely suppressed in Pt<sub>4</sub>Ge<sub>3</sub> under the conditions studied in this work (see Figure 6b). CO coverage is negligible even at P<sub>CO</sub> > P<sub>O<sub>2</sub></sub>. Our results indicate that the bistable kinetics of CO oxidation disappears; a single regime of fast CO oxidation rate is obtained by controlling the promoter concentration. The three different pressure conditions show very similar behaviour.

$O_2^*$  and  $O^*$  are always the only species found at steady-state. The  $O_2^*$  coverage changes in a pyramid-like fashion, showing the highest value at intermediate temperatures. At  $P_{O_2}=P_{CO}$ ,  $O_2^*$  remains the main dominant species until 350 K, whereas at  $P_{O_2}>P_{CO}$ ,  $O^*$  coverage becomes larger in the whole temperature range due to greater availability for  $O_2$  dissociation sites. We highlight that the Pt–Ge nanoalloy starts to show very similar  $BGs$  for CO and  $O_2$  at this precise Ge concentration. Thus, both reactants will adsorb with similar probabilities, and according to the Sabatier principle, the conditions for CO oxidation will be optimal. The importance of the magnitude of the binding energies on the overall performance of the reaction kinetics has been underlined previously.<sup>[69]</sup> In addition,  $Pt_4Ge_3$  shows much faster kinetics for the  $CO^*+O^*$  than  $CO^*+O_2^*$  step (see Table S3 for the corresponding reaction rates). This means that at initial times the  $TOF_{CO_2}$  will be small as there will only be present  $CO^*$  and  $O_2^*$ . But once  $O_2$  starts dissociating,  $CO^*$  will react immediately with the newly formed  $O^*$  to give  $CO_2$ .  $O_2^*$  accumulation is not problematic since it can dissociate to  $O^*$  or react with CO(g) via ER. It is also important to mention that  $CO_2$  desorption becomes spontaneous with increasing temperature, which prevents  $CO_2$  poisoning.

Except at  $P_{O_2}>P_{CO}$ , the  $TOF_{CO_2}$  is higher at low temperatures and slightly decreases monotonically with temperature, but always maintains at high value. Interestingly, the highest  $TOF_{CO_2}$  is obtained at 100 K for  $P_{O_2}=P_{CO}$  and  $P_{CO}>P_{O_2}$ , which is at least one order of magnitude higher as compared to other compositions. Part of these results may be found somewhat counterintuitive. However, we believe that this behaviour might be the result of the interplay between two effects, *i.e.*, entropic and the dynamic nature of the steady-state. First, the DFT results reflect that the Gibbs free energies of the reactions and barriers have a substantial dependence with temperature (see for instance Figure S11). This alters the viability of the different reaction channels (e.g.,  $CO^*+O^*$  vs  $CO^*+O_2^*$ ) at different temperatures, and thus, yields different steady-state coverages and TOFs. Second, it should be kept in mind that in CSTR conditions the system is not at thermodynamic steady-state (equilibrium) but at a dynamic steady-state, which depends on the gas in/out flow. Namely, the constant flow of gas species also affects the surface composition and can in turn shift the elementary reactions at either direction with different strength. One should be cautious to rely solely on equilibrium data to rationalize the behaviour of systems out of thermodynamic equilibrium. The changes occur so that  $O_2^*$  is formed at the expense of  $O^*$ , showing a pyramid-like behaviour. Unfortunately, quantifying this hypothesis is out of the scope of this work. Preliminary steps were done by performing a degree of rate control (DRC) analysis, but unfortunately we faced numerical instabilities with the evaluation of numerical derivatives. Finite difference method has already been reported problematic previously,<sup>[70]</sup> especially in transient kinetics where the magnitude of the finite-size perturbation can affect the convergence of the dynamic steady-state.<sup>[71]</sup>

Besides, it is important to point out that  $Pt_4Ge_3$  will also get poisoned eventually at higher CO pressures. The bistable kinetics is a property inherent to CO oxidation. Ge alloying acts

by increasing the boundaries of the fast kinetic regime to harsher conditions in the P–T phase diagram. The CO-poisoning region does not disappear, although it will occupy a much smaller portion of the phase diagram compared with pure  $Pt_4$ .

Regarding the role of Ge, DFT results suggest that on the one hand, Ge induces ligand effects (electronic effects) on Pt and reduces the affinity to CO, avoiding the overbinding of CO over  $O_2$ . Similar ligand effects were also observed by Molina *et al.* for Nb and Mo doping.<sup>[72]</sup> The advantageous mixing between both elements guarantees that Ge hybridizes with most Pt sites in the nanoalloy. On the other hand, our theoretical findings suggest that Pt–Ge also acts as a bifunctional catalyst. The addition of Ge into the catalyst creates new active sites for a more efficient oxygen adsorption. The nanoalloys show a dual adsorption behavior where CO is preferentially adsorbed on Pt and  $O_2$  on Ge avoiding the competition for the active sites. Such behavior is in line with other oxophilic promoters such as Ru, Mo or Sn.<sup>[18,26]</sup> Although in many instances Pt–Ge ensembles are needed for  $O_2^*$  and  $O^*$  bridge adsorption, the ligand effects provide a similar adsorption strength to CO. Again, the high mixing guarantees that  $CO^*$  and  $O_2^*$  (or  $O^*$ ) sites are in proximity for a facile diffusion and reaction, maximizing the bifunctional effect. Importantly, the ligand and bifunctional effects depend on Ge content, *i.e.*, Ge alloying offers a new degree of freedom to control the catalytic properties of bimetallic catalysts. Apparently, this concentration dependence is very reaction-specific. Our results suggest that  $Pt_4Ge_3$  composition has the most adequate reactivity for CO oxidation, in contrast to dehydrogenation reactions where  $Pt_4Ge$  outperforms  $Pt_4Ge_3$ .<sup>[33]</sup>

The microkinetic simulations further confirm that Ge alloying alters the CO oxidation bistable kinetics. Ge leads to a CO poisoning-free regime and fast  $CO_2$  production prevails, at least within the pressure ranges studied in this work. Besides, the microkinetic simulations reveal that in this regime the Pt–Ge clusters may exhibit a high oxidation state (high  $O^*$  coverage). Although our microkinetic model does not distinguish between Pt and Ge sites, the information about the binding energies is included; and thus, this high oxidation state may be indicative of a bifunctional mechanism. By virtue of these findings and previous work on bigger NPs,<sup>[40]</sup> we speculate that part of the subnano catalyst may remain reduced to the metallic state with Pt–Ge alloy formation, and part may be oxidized to  $GeO_x$ . Therefore, now there is available an independent oxygen reservoir resistant to CO poisoning, *i.e.*,  $GeO_x$ , which can react with CO on Pt metallic sites.

## Conclusions

We have studied the effect of the composition on the CO oxidation reaction catalyzed by  $Pt_nGe_m$  clusters ( $n=0-4$ ) supported on MgO(100) using a combination of DFT and microkinetic simulations. To characterize the most stable CO and  $O_2$  adsorption configurations, a global minima search was carried out on the thermally populated isomers of each cluster composition. The DFT results reveal that CO binding energy

weakens with the Ge content owing to ligand (electronic) effects induced on Pt upon Ge alloying. Bond-length and PDOS analysis indicate that O<sub>2</sub> is readily activated up to the peroxy state and dissociates to atomic oxygen with low barriers and significant release of energy in all the Pt<sub>4</sub>Ge<sub>n</sub> clusters.

A systematic study of CO oxidation is performed by considering the LH and ER mechanisms initiated by both molecular and atomic oxygen. In the ER mechanism pure and alloyed clusters present similar barriers and the whole process is downhill. Nonetheless, a pronounced dependence on the Ge concentration is found for the LH pathway, which is the most likely reaction channel. For Pt<sub>4</sub>, the barriers for the CO\* + 2×O\* are ca. 1 eV. In addition, pure Pt is prone to become poisoned by CO and CO<sub>2</sub>. On the contrary, the poisoning effects are overcome on Pt–Ge. Pt<sub>4</sub>Ge, Pt<sub>4</sub>Ge<sub>2</sub> and Pt<sub>4</sub>Ge<sub>4</sub> present acceptable reaction barriers except one of the steps in the CO\* + 2×O\* route which is blocked with relatively high activation and reaction energies. For Pt<sub>4</sub>Ge<sub>3</sub>, all the reaction barriers are attainable (< 1 eV).

The microkinetic simulations hint to a reduced CO and CO<sub>2</sub> poisoning dependence with increased Ge concentration. Pt<sub>4</sub> shows a bistable-like CO oxidation kinetics, in agreement with the literature. At low temperatures and/or high CO pressures it suffers from CO poisoning and CO<sub>2</sub> production drops drastically. On top of this, regardless of partial pressures, there is a constant CO<sub>2</sub> poisoning which also diminishes the TOF<sub>CO<sub>2</sub></sub>. Nevertheless, as the Ge concentration increases CO and CO<sub>2</sub> poisoning subside, until in Pt<sub>4</sub>Ge<sub>3</sub> the kinetic regime of severe CO poisoning is removed. A 4:3 Pt to Ge ratio changes the bistable kinetics of pure Pt and instead, a regime of fast CO<sub>2</sub> formation is obtained for all temperatures and pressures considered. CO<sub>2</sub> poisoning does not change in the same way with the Ge content as Pt<sub>4</sub>Ge<sub>4</sub> again shows some CO<sub>2</sub> poisoning. Thus, Pt<sub>4</sub>Ge<sub>3</sub> appears to be the optimal composition for the CO oxidation reaction.

Our calculations predict that Ge alloying induces both ligand and bifunctional effects. Charge transfer phenomena and orbital hybridization between Pt and Ge alters the adsorption properties of the catalyst (structure-activity relationship) in a composition dependent fashion. Further, Pt–Ge shows a dual adsorption behavior where CO always binds to Pt sites, while oxygen shows a clear preference towards Ge due to its greater oxophilicity. Precisely, this oxophilicity of Ge makes part of it oxidize to GeO<sub>x</sub> and consequently, a new oxygen source is developed resistant to CO poisoning. The provided results require a more thorough analysis to correctly sample the boundary between the CO poisoning and CO poisoning-free regimes as a function of pressures and temperature. As for now, this remains yet a topic of future study.

On the basis of this work, we believe that Ge alloying can be used for tuning the properties of Pt for catalytic applications. The high CO poisoning resistance conferred by Ge alloying without harming the catalytic activity makes it very appealing for industrial applications such as the water-gas shift reaction and hydrogen fuel cells.

## Supporting Information

The authors have cited additional references within the Supporting Information (Ref. [73,74]).

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## Conflict of Interests

The authors declare no competing financial interest.

## Data Availability Statement

The primary data that support the findings of this study are available within the article; additional data are available from the corresponding author upon reasonable request.

**Keywords:** bifunctional catalyst · CO oxidation · CO poisoning · Ge alloying · microkinetic modeling

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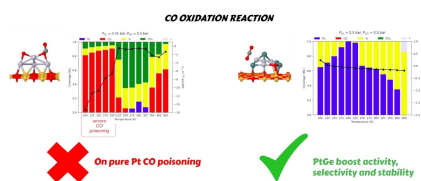
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# RESEARCH ARTICLE

**PtGe nanoalloys** act as a bifunctional catalyst where Ge forms an alternative oxygen source  $\text{GeO}_x$  and modifies the electronic structure of Pt (ligand effects) and achieves a CO poisoning-free kinetic regime.



*A. Ugartemendia, Prof. J. M. Mercero\*, Prof. A. de Cózar, Dr. M. M. Melander, Prof. J. Akola\*, Prof. E. Jimenez-Izal\**

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**Deposited PtGe Clusters as Active and Durable Catalysts for CO Oxidation**

